School of Basic Sciences and Research

Department of Chemistry and Biochemistry

Programme and Course Structure
AY: 2020-22

MSc. in Chemistry

Program Code: SBR0101

1.1 Vision, Mission and Core Values of the University

Vision of the University

To serve the society by being a global University of higher learning in pursuit of academic excellence, innovation and nurturing entrepreneurship.

Mission of the University

1. **Transformative educational experience**
2. **Enrichment by educational initiatives that encourage global outlook**
3. **Develop research, support disruptive innovations and accelerate entrepreneurship**
4. **Seeking beyond boundaries**

Core Values

- **Integrity**
- **Leadership**
- **Diversity**
- **Community**

1.2 Vision and Mission of the School

Vision of the School

Achieving excellence in the realm of science to address the challenges of evolving society

Mission of the School

1. **To equip the students with knowledge and skills in basic and applied sciences**
2. **Capacity building through advanced training and academic flexibility.**
3. **To establish center of excellence for ecologically and socially innovative research.**
4. **To strengthen interinstitutional and industrial collaboration for skill development and global employability.**

1.3 Vision and Mission of Department of Chemistry & Biochemistry

Vision of Chemistry & Biochemistry

Strive to achieve excellence in teaching and research in the field of Chemistry and Biochemistry and to build human resource for solving contemporary problems.

Mission of Chemistry & Biochemistry

- **Providing distinctive and relevant education in Chemistry and Biochemistry to students.**
- **Motivating young minds through innovative teaching methods, to acquire theoretical knowledge and practical skills in different disciplines of chemistry and empowering them with problem solving skills.**
- **Nurturing innovation by carrying out world class research and scholarly work**
- **Promoting interdisciplinary research in collaboration with national/international laboratories/Institutions.**

1.3 Programme Educational Objectives (PEO)

Program educational objectives are broad statements that describe the career and professional accomplishments that the program is preparing Post graduates to achieve.

PEO 1: To prepare students for advanced studies in Chemistry and its allied fields.

PEO2: To help students understand the value of advanced characterization techniques by gaining the knowledge of spectroscopy, chromatography and chemistry of natural products.

PEO 3: To expose the students to the practical aspects of chemistry by means of qualitative, quantitative and advance instrumental methods.

PEO 4: To develop the ability to communicate scientific information in written and oral formats.

1.3.3 Program Outcomes (PO's)

PO1: Gained knowledge, abilities and insight in well defined area of research within Chemistry.

PO2: Competency to work effectively and safely in a laboratory environment.

PO3: Developed communication skills, both written and oral, for specific for specialized audiences.

PO4: Acquired the skills of planning and conducting advanced chemical experiments and applying structural-chemical characterization techniques.

PSO1 : Global level research opportunities to pursue Ph.D. programme and target the CSIR – NET examination.

PSO2 : Explain the properties of metal Complexes and Transition Elements.

PSO3 : Use spectrophotometer to find out the allowedness and analysis of known and unknown organic, inorganic and natural compounds through UV, IR NMR and Mass data.

PSO4 : The broad education necessary to understand the impact of chemical solutions in a global and societal context.

Program Structure
School of Basic Sciences & Research
M. Sc. Chemistry
Batch: 2020-22
TERM: I

S. No.	Subject Code	Subjects	Teaching Load			Credits	Pre-Requisite/Co Requisite
			L	T	P		
THEORY SUBJECTS							
1.	MCH131	Inorganic Chemistry-I	4	0	0	4	Core
2.	MCH132	Organic Chemistry-I	4	0	0	4	Core
3.	MCH133	Physical Chemistry-I	4	0	0	4	Core
4.	MCH134	Analytical Chemistry-I	4	0	0	4	Core
5.	MMT129	Introduction to MATLAB & its application	3	0	0	3	GE
PRACTICAL							
6.	MCH171	Inorganic Chemistry Lab-I	0	0	3	2	Core
7.	MCH172	Organic Chemistry Lab-I	0	0	3	2	Core
8.	MCH173	Physical Chemistry Lab-I	0	0	3	2	Core
TOTAL CREDITS						**25**	

Program Structure
School of Basic Sciences & Research
M. Sc. Chemistry
Batch: 2020-2022
TERM: II

S. No.	Course Code	Course	Teaching Load			Credits	Core/Elective
			L	T	P		
THEORY SUBJECTS							
1.	MCH135	Inorganic Chemistry-II	4	0	0	4	Core
2.	MCH136	Organic Chemistry-II	4	0	0	4	Core
3.	MCH137	Physical Chemistry-II	4	0	0	4	Core
4.	MCH138	Analytical Chemistry-II	4	0	0	4	Core
5.	MPH115	Renewable Energy Sources: Solar And Hydrogen Energy	4	0	0	4	GE
6.	CCU401	Community Connect	2	0	0	2	SEEC-1
PRACTICAL							
7.	MCH174	Inorganic Chemistry Lab-II	0	0	3	2	Core
8.	MCH175	Organic Chemistry Lab-II	0	0	3	2	Core
9.	MCH176	Physical Chemistry Lab-II	0	0	3	2	Core
TOTAL CREDITS						**28**	

Program Structure Template
School of Basic Sciences & Research
M. Sc. Chemistry
Batch: 2020-2022
TERM: III

S. No.	Course Code	Course	Teaching Load			Credits	Core/Elective
			L	T	P		
THEORY SUBJECTS							
1.	MCH231	Molecular Spectroscopy	4	0	0	4	Core
2.	MCH232/MCH233/MCH234	Inorganic Chemistry-III/ Physical Chemistry-III/ Organic Chemistry-III	4	0	0	4	Core
3.	MCH235/MCH236/MCH237	Inorganic Chemistry-IV/ Physical Chemistry-IV/ Organic Chemistry-IV	4	0	0	4	Core
4.	MCE201/202	Environmental Chemistry / Polymer Science and Technology	4	0	0	4	DSE
PRACTICAL							
5.	MCH271/ MCH272/273	Organic Chemistry Lab-III/ Physical Chemistry Lab-III/ Inorganic Chemistry Lab-III	0	0	4	2	Core
6.	MCH276	Dissertation-Part-A	0	0	6	2	Core
TOTAL CREDITS						**20**	

Program Structure Template
School of Basic Sciences & Research
M. Sc. Chemistry
Batch: 2020-2022
TERM: IV

S. No.	Course Code	Course	Teaching Load			Credits	Core/
			L	T	P		
THEORY SUBJECTS							
1.	MCH238/MCH239/MCH240	Inorganic Chemistry-V/ Physical Chemistry-V/ Organic Chemistry-V	4	0	0	4	Core
2.	MCH241/MCH242/MCH243	Inorganic Chemistry-VI/ Physical Chemistry-VI/ Organic Chemistry-VI	4	0	0	4	Core
3.	MCE203/204	Medicinal Chemistry/ Chemistry of Nanomaterials	4	0	0	4	DSE
4.	OPEXXX	Open Elective	2	0	0	2	SEEC-2
Practical							
5.	MCH275	Dissertation-Part-B	0	0	12	6	Core
TOTAL CREDITS						**20**	

C. Course

- **Theory Subject**
- **Practical Subjects**
- **Projects/Dissertations**

2.1 Template A1: Inorganic Chemistry-I (MCH131)

School: SBSR		**Batch 2020-22**
Program: M.Sc.		**Current Academic Year : 2020-22**
Branch:Chemistry		**Semester I**
1	**Course Code**	**MCH131**
2	**Course Title**	**Inorganic Chemistry I**
3	Credits	4
4	Contact hours	4-0-0
	Course Status	Compulsory
5	Course Objectives	1.To provide an insight into bonding and structure of coordination compounds. 2.To explain the spectral and magnetic behaviour of coordination compounds. 3.To provide a thorough knowledge about the chemistry and application of inner transition metals. 4.To discuss about various spectroscopic methods for structure elucidation of inorganic compounds. 5.To explain the basics of radioactivity as well as various radio analytical techniques. 6.To impart knowledge about structure, bonding and application of inorganic compounds and radio chemistry.
36	Course Outcome	CO1 : Explain the various theories of metal –ligand bonding CO2 : Explain the electronic spectra and magnetic properties of transition metal complexes. CO3 : Interpret the EPR and Mossbauer spectra CO4 : Illustrate the chemistry and uses of inner transition metals CO5 : Know about various radio-analytical techniques CO6 : Gain knowledge about of various aspects of modern inorganic chemistry
7	Course Description	This course include basic concepts of metal –ligand bonding, magnetic and electronic properties of coordination compounds and their characterization techniques. Chemistry of inner transition metals and nuclear chemistry are also discussed in this course.
8	Outline Syllabus	
	Unit 1	**Metal-ligand Bonding**
	A	Overview of crystal field and ligand field theories of 4-, 5-and 6-coordinated complexes, d-orbitals splitting in linear, trigonal, octahedral, square planar, tetrahedral, square pyramidal, trigonal-bipyramidal and cubic complexes
	B	measurement of CFSE (d^1 to d^{10}) in weak and strong ligand fields, JahnTeller distortion, nephelauxetic series
	C	Molecular orbital theory (MOT) of coordination compounds: Composition of ligand group orbitals, molecular orbital energy diagrams of octahedral, tetrahedral, square planar complexes including both s and p

		bonding, angular overlap model		
	Unit 2	**Electronic Spectra and Magnetic Properties of Transition Metal Complexes**		
	A	Interpretation of electronic spectra, Orgel diagrams, Tanabe-Sugano diagrams for transition metal complexes (d^1 - d^9 states), calculations of Dq, B and β parameters		
	B	charge transfer spectra, spectroscopic method of assignment of absolute configuration in optically active metal chelates and their stereochemical information		
	C	anomalous magnetic moments, magnetic exchange coupling, temperature independent paramagnetism (TIP) of complexes, spin cross over phenomenon. Effect of temperature on their magnetic properties		
	Unit 3	**Chemistry of Inner Transition Elements**		
	A	General discussion on the properties of the f-block elements.		
	B	Redox, Spectral and Magnetic properties.		
	C	Use of Lanthanide compounds as shift reagents. Photophysical properties of Lanthanide complexes.		
	Unit 4	**Characterization Techniques**		
	A	EPR spectroscopy-basic principle, hyperfine and superhyperfine lines, anisotropy, g values, application in selected inorganic compounds.		
	B	Mossbauer Spectroscopy-Gamma ray emission and absorption by nuclei, Mossbauer effect — conditions, Doppler effect, instrumentation, chemical shift examples, quadrupole effect,		
	C	Use of Mössbauer spectra in chemical analysis, typical spectra of iron and tin compounds.Optical rotatory dispersion (ORD) and circular dichroism (CD).		
	Unit 5	**Nuclear Chemistry**		
	A	Nuclear structures and nuclear stability. Nuclear models ; radioactivity and nuclear reactions. Detection and measurement of radiation. Tracer techniques.		
	B	Study of chemical reactions, isotope exchange reactions, kinetic isotope effect, nuclear activation analyses, Principle of nuclear detection, gas detector, ionization chamber, proportional and G. M. detector.		
	C	*Radioactive Techniques:* Detection and measurement of radiation- GM ionization and proportional counters. Radiometric analysis: Isotope dilution analysis, age determination, neutron activation analysis (NAA) and their applications. Radiation hazards and safety measures.		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1.Inorganic Chemistry, J.E. Huhey, Harper & Row.		
	Other References	1.Concise Inorganic Chemistry,J. D. Lee, Elbs with Chapman and Hall, London. 2.The Chemical bond, J.N.Murre l, SFA Kettle and JM. Tedder, Wiley, New York. Advanced Inorganic Chemistry, F.A. Cotton and Wilkinson, John Wiley.		

2.1 Template A1: Organic Chemistry-I (MCH132)

School: SBSR		**Batch 2020-22**
Program: M.Sc.		**Current Academic Year : 2020-22**
Branch : Chemistry		**Semester I**
1	**Course No.**	**MCH132**
2	**Course Title**	**Organic Chemistry 1**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course status	Compulsory
5	Course Objective	1.To enhance the analytical ability of students about the basic and modern concepts of conjugation, resonance and aromaticity. 2. To impart knowledge of mechanistic, kinetic and thermodynamic aspects of i. nucleophilic and electrophilic substitution. ii. Reaction conditions, products formation and mechanisms of some named reactions. iii. addition reactions of C=C and C=O bonds and elimination reactions. 3.To teach the concepts and critical bond forming reactions and reaction intermediates in organic synthesis and molecular rearrangements 4. To make the student conversant with - the basic concepts in stereochemistry. 5.To discuss the Conformational analysis, reactivity, chirality, interconversion, resolution and asymmetric synthesis.
6.	Course Outcomes	The students will acquire the knowledge and analytical ability to CO1. Rationalize the concept of Aromaticity, nonaromaticity and antiaromaticity in carbocyclic and heterocyclic compounds CO 2. Solve the reactions and analyze the conditions, products formation and mechanisms of different reactions. CO3. Recognize the correct reaction intermediate formation and different aspects of their stability and reactivity. CO4. Critically examine the chirality/prochirality in the molecules and understand the enentio and diastereospecific/selective reactions. CO5. Conformational analysis of cycloalkanes, reactivity, chirality, interconversion, resolution and asymmetric synthesis,. CO6. The students will be able to acquire the skills for recognizing the reaction, rationalizing their mechanism, kinetic vs. thermodynamic considerations critical analysis of intermediates and correct stereochemical assignment and interpretation in rather simple organic molecules and reactions.
7	Course Description	This course tends to the build the foundation of Organic Chemistry that provides the insight of different aspects of organic reactions in terms of energy considerations, stereochemical implications, mechanistic approach and involvement of intermediates.

8	Outline syllabus	
	Unit 1	**Nature of Bonding in Organic Molecules**
	A	Delocalized chemical bonding: conjugation, cross conjugation, resonance, hyperconjugation, tautomerism;
	B	Criteria for aromaticity: Huckel's 4n+2 electron rule for benzenoid and non benzenoid aromatic compounds; Application in carbocyclic and heterocyclic systems, n-annulenes, heteroannulene, fullerenes, C-60, cryptates, azulenes.
	C	Current concepts of aromaticity: Anti-aromatic, non-aromatic and homoaromatic compounds, Effect of tautomerism and hyperconjugation on aromaticity.
	Unit 2	**Reaction Mechanism - Structure and Reactivity**
	A	Types of reaction mechanisms- substitutions, eliminations, additions, rearrangements, thermodynamic and kinetic requirements
	B	Hammond postulate, Curtin-Hammett principle, transition states and intermediates, catalysis: electrophilic catalysis, acid and base catalysis
	C	Libido rule; methods of determination of reaction mechanism methods of determining mechanisms, isotopic effects.
	Unit 3	**Reaction Intermediates**
	A	Classical and non classical carbocations, phenonium ions, norbornyl system, common carbocation rearrangement (Wagner Meerwein rearrangement, Demjonove rearrangement and Pinacol-pinacolone rearrangement);
	B	Carbanions: ambident ions and their reactions. HSAB principle and its applications;
	C	Free radicals: cage effects. Radical Cations and Radical Anions; Carbene: Synthesis, structure and reactions of singlet and triplet carbene, nitrenes, Benzyne.
	Unit 4	**Stereochemistry I**
	A	Elements of symmetry, chirality (centre, axis and plane), molecules with more than one chiral center, threo and erythro isomers, optical purity
	B	Topicity of ligand and faces and their nomenclature, stereogenecity, chirogenicity and pseudosymmetry, stereospecific and stereoselective reactions
	C	Asymmetric synthesis: Chiral auxiliaries, methods of asymmetric induction – substrate, reagent and catalyst controlled reactions; determination of enantiomeric and diastereomeric excess; enantio-discrimination. Resolution – optical and kinetic
	Unit 5	**Stereochemistry II**
	A	Conformational analysis of cyclic systems: Cyclohexane and its derivatives (mono-,and di- substituted), fused (decalins) and bridged bicyclic systems, effect of conformation on the reduction of cyclic ketones,
	B	nucleophilic addition to carbonyl group (Cram, Franklin Ahn Model, Cieplak effect), nucleophilic substitution on cyclohexane substrates, cyclohexane epoxide formation and opening

	C	elimination reactions of cyclohexyl halides, de-amination of 2-aminocyclohexanols, elimination vs substitution competition and neighboring group participation reactions of acyclic and cyclic molecules.		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text Book	1.Stereochemistry, P. S. Kalsi, New Age International. 2.Organic Chemistry, R. T. Morrison and R. N. Boyd, Prentice-Hall. 3. Reaction Mechanism in Organic Chemistry, S. M. Mukherji and S. P. Singh, Macmillan.		
	Other references	1. Advanced Organic Chemistry Reactions, Mechanism and Structure, Jerry March, John Wiley. 2.Stereochemistry of Organic Compounds By Ernest Ludwig Eleil, Samual H. Wilen. 3.Stereochemistry of Organic Compounds: Principles and Applications by D. Nasipuri		

2.1 Template A1: Physical Chemistry-I (MCH133)

School: SBSR		**Batch : 2020-22**
Program:M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester:I**
1	**Course Code**	**MCH133**
2	**Course Title**	**Physical Chemistry I**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objective	1. To provide the understanding of physical states of matter and their practical applications. To define how the initially primitive models of real gases in physical chemistry are elaborated to take into account more detailed observations. 2. To understand the concept of partial molar quantities and their variation with temperature and pressure. 3. The concept of ensembles, partition function and their applications in studying gaseous molecules. 4. To understand the concept and different theories of ions and electrolyte interactions 5.To discuss the theoretical aspects of chemical kinetics and the importance of rate equations and different theories for studying the kinetics of complex reactions. 6. To provide an in-depth analysis of various phenomenon, laws and applications of States of Matter, Thermodynamics, Electrochemistry, Phase Equilibrium and Chemical Kinetics
6	Course Outcomes	CO1: Understand the detailed concept of liquid and gaseous state and the structural features of solid state material by having complete knowledge of X-ray diffraction and its analysis. CO2: Understand the application of second law of thermodynamics and the concept of third law of thermodynamics. CO3: Familiarize with the applications of partition function and statistics in understanding the thermodynamics of molecules. CO4: Understand the concept of electrical double layer at the electrode electrolyte interface by studying different proposed models of it. CO5: Understand the detailed concepts of kinetics and its applications, Influence of physical and chemical parameters on reaction rates in solutions CO6: Do the in-depth analysis of various phenomenon and laws of States of Matter, applications of Thermodynamics, Electrochemistry and Chemical Kinetics and different functions of statistical thermodynamics.
7	Course Description	The course is framed to give broad view of states of matter, chemical potential, concepts of electrical double layer in solutions and various models to explain it. Concept of existence of different phases in the form of phase diagrams and their existence with changing variables.

8	Outline syllabus	
	Unit 1	**States of Matter**
	A	(a) Gaseous State : Maxwell–Boltzmann distribution of molecular velocities of gases (b) Liquid State: Structure of liquids, Radial distribution functions
	B	Monte–Carlo method, Molecular dynamics.(c) Solid State: Types of solids, Debye- Scherrer method of X-ray structure analysis of crystals, indexing of reflections,
	C	structure of simple lattice and X-Ray intensities, structure factor and its relation to intensity and electron density, Rietveld analysis, particle size of crystallites.
	Unit 2	**Thermodynamics**
	A	Essentials of thermodynamics, fugacity, standard state of real gases, the relation between fugacity and pressure, Partial molar quantities, chemical potential and Gibbs-Duhem equation,
	B	Classius – Clayperon equation; law of mass action and its thermodynamic derivation, variation of chemical potential with temperature and pressure, chemical potential for an ideal gas, determination of partial molar volume
	C	thermodynamic functions of mixing (free energy, entropy, volume and enthalpy), third law of thermodynamics, residual entropy, meaning and scope of irreversible thermodynamics.
	Unit 3	**Statistical Thermodynamics**
	A	Concept of distribution, Thermodynamic probability and most probable distribution. Ensembles, Canonical, grand canonical and microcanonical ensembles.
	B	Partition function - Translational, Rotational, Vibrational and Electronic partition functions, calculation of thermodynamic properties in terms of partition function. Applications of partition functions.
	C	Heat capacity behaviour of solids - Chemical equilibria and equilibrium constant in terms of partition functions, Fermi-Dirac statistics, distribution law, Bose-Einstein statistics - distribution law, Evaluation of Lagrange's undetermined multipliers.
	Unit 4	**Electrochemistry**
	A	Debye-Huckel theory of ion- ion interactions, Debye-Huckel limiting law of activity coefficients and its limitations,
	B	Debye - Huckel -Onsager treatment for aqueous solutions and its limitations, Wein effect, Debye – Falkenhagen effect.
	C	The electrode-electrolyte interface: The electrical double layer -The Helmholtz-Perrin parallel plate model, the Gouy-Chapman diffuse-charge model and the Stern model, excess function
	Unit 5	**Chemical Kinetics**
	A	Simple collision theory of reaction rates, Arrhenius equation and activated complex theory (ACT), thermodynamic treatment, chain reactions (hydrogen-halogen reactions) decomposition of N_2O_5
	B	Theory of unimolecular reactions: Lindemann – Hinshelwood

		mechanism of unimolecular reactions, RRKM and Slater treatment,		
	C	Factors affecting rate of chemical reactions in solution Effect of solvent and ionic strength (Primary salt effect) on rate constants, secondary salt effect.		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1.Physical Chemistry, P. W. Atkins, Oxford University Press, New York. 2.Textbook of Physical Chemistry by K. L. Kapoor (Volume 1) 3.Textbook of Physical Chemistry by K. L. Kapoor (Volume 3) 4.Textbook of Physical Chemistry by K. L. Kapoor (Volume 5)		
	Other References	1. Physical Chemistry, I.N. Levine, Tata McGraw Hill Pub. Co. Ltd., New Delhi. 2. Comprehensive Physical Chemistry by N.B.Singh, N.S.Gajbhiye and S.S.Das , New Age publishers, New Delhi 3. Chemical Kinetics, K. J. Laidler, Harper & Row, New York. 4. Physical Chemistry by D.A.McQuarrie and J.D.Simon		

2.1 Template A1: Analytical Chemistry-I (MCH134)

School: SBSR		**Batch : 2020-2022**
Program: M.Sc		**Current Academic Year: 2020-22**
Branch: Chemistry		**Semester: I**
1	**Course Code**	**MCH134**
2	**Course Title**	**Analytical Chemistry I**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objective	1.Provide and enrich the students to analytical techniques, various types of errors knowingly/ unknowingly introduced, accuracy and confidence limit in analytical process. 2.Provide detailed insight of chemical equilibrium and its effect in chemical analysis of analyte. 3.Provide detailed technical knowledge of various chromatogaraphic separation techniques based on physical state, contact and separation mechanism. 4.Provide detailed technical knowledge of gas, thin layer chromatographic, integrated LC-MS and GC-MS separation techniques for qualitative and quantitative analysis. 5.Enable the students to study the thermal behaviour of different compounds and study temperature dependent decomposition process and structural elucidation of unknown analyte. CO6:Estimate the temperature dependent weight loss in compound and model and optimize suitable temperature condition for further chemical processing.
6	Course Outcomes	CO1: Apply the knowledge of analytical techniques to minimize the error and report the outcomes of analysis with high precision and accuracy, CO2: Understand the role of different analytical techniques used for the separation of compounds present in very small quantity, CO3:Understand the role of chemical equilibrium in chemical analysis, CO4: Segregate and select the suitable indicator for measurement of pH, CO5: Purify the various compounds for their further detailed structural elucidation and molecular mass analysis, CO6. To learn analytical tools involving Chromatographic methods and thermo-analytical instruments of a lab for the identification of equilibrium process.
7	Course Description	Analytical chemistry I emphasizes on various factors as - types of errors, accuracy and precision in chemical analysis, concepts of chemical equilibrium and its effects on qualitative and quantitative estimation, Chromatographic separation and Thermal analysis.
8	Outline syllabus	
	Unit 1	**Introduction to Analytical Chemistry**

	A	Scope & objectives of Analytical chemistry and chemical analysis, Classification of analytical methods. Errors in chemical analyses- Accuracy and precision
	B	Types of error-determinant, indeterminate and gross. Nature of random errors, statistical treatment of random errors, standard deviation of calculated results, reporting of calculated data
	C	ways of expressing accuracy and precision. variance and confidence limit. Comparison of mean with true values, regression analysis (least-square method for linear plots)
	Unit 2	**Concept of Equilibrium**
	A	General treatment of equilibria in aqueous medium involving monoprotic weak acid and weak base, and salts of weak acids and weak bases
	B	Activity and activity coefficient; Effect of electrolytes on chemical equilibria, Calculation of pH
	C	Constructing titration curves from charge balance and mass balance equations, Acid-base titrations and theory of pH indicators.
	Unit 3	**Chromatographic Methods-I**
	A	General principle, classification of chromatographic methods based on physical state, contact and separation mechanism
	B	Nature of partition forces. Chromatographic behavior of solutes. Chromatographic resolution, selectivity factor and column efficiency.
	C	Column chromatography**:** Nature of column materials, Preparation of the column, Solvent systems, detection methods and applications.
	Unit 4	**Chromatographic Methods-II**
	A	**Gas chromatography-** principle, experimental technique, carrier gas, sample injection, column, detector and application
	B	High Performance Liquid Chromatography (HPLC): instrumentation-solvent and reservoirs, pumping system, sample injection, Column, detectors
	C	**Thin layer chromatography:** coating of materials, preparation of TLC, Solvents, methods of detection and applications. Theory and application of LC-MS, Pyrolysis GC-MS, Thermal Desorption GC-MS.
	Unit 5	**Thermal Analysis**
	A	Principle, different methods of thermal analysis, i) Thermo gravimetric methods of analysis(TG/DTG): Instrumentation, thermogram and information from thermogram, factors affecting thermogram, applications TGA for quantitative analysis (TG analysis of $CaC_2O_4.H_2O$, $CuSO_4.5H_2O$, dolomite ore, etc.)
	B	Problems based TGA, ii) Differential Thermal Analysis (DTA): Instrumentation, general principles, differential thermogram, DTA and TG curve together, Applications (DTA analysis of mixture of polymers, DTA of CaC_2O_4 H_2O, DTA of $CuSO_4$ $5H_2O$).
	C	Differential Scanning Calorimetry (DSC): Principle, Instrumentation, and Applications (DSC curve of polyethylene terphthalate, DSC curve for isothermal crystallization of polyethylene, DSC of phenacetein), thermometric titrations, Evolved gas analysis.

	Mode of examination	Theory		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1.Analytical Chemistry-An Introduction, 7th Edition,D. A. Skoog, D.M. West, F.J. Holler, S.R. Crouch, Saunders College Publishing, Philadelphia, London.		
	Other References	1. Modern Methods of Chemical Analysis, 2nd Edition,R. L. Pecsok, L. D. Shields, T. Cairns and L.C. Mc William, John Wiley, New York. 2. Analytical Chemistry, 5th Edition,G. D. Christian, John Wiley & Sons, New York. 3. Analytical Chemistry: Principles, 2nd Edition,J. H. Kennedy, Saunders Holt, London.		

2.1 Template A1: INTRODUCTION TO MATLAB AND ITS APPLICATIONS (MMT129)

School: SBSR		**Batch : 2020-2022**
Program: M.Sc.		**Current Academic Year: 2020-22**
Branch: Mathematics		**Semester: I**
1	Course Code	MMT-129
2	**Course Title**	**INTRODUCTION TO MATLAB AND ITS APPLICATIONS**
3	Credits	**3**
4	Contact Hours (L-T-P)	3-0-0
	Course Status	Compulsory
5	Course Objective	The goal of this course is to introduce the necessary mathematical concepts for MATLAB and cover the syntax and semantics of MATLAB including control structures, comments, variables, functions etc. Once the foundations of the language have been established students will explore different types of scientific programming problems including curve fitting, ODE solving etc.
6	Course Outcomes	• CO1: Describe the fundamentals of MATLAB and use MATLAB for interactive computations. (K2, K3) • CO2: Demonstrate with strings and matrices and their uses. (K2, K3) • CO3: Illustrate basic flow controls (if-else, for, while). (K3) • CO4: Create plots and export this for use in reports and presentations. (K3, K5) • CO5: Develop program scripts and functions using the MATLAB development environment. (K4, K5) • CO6: Write the program for evaluates linear system of equations, ordinary differential equations in MATLAB. (K5,K6)
7	Course Description	The course will give the fundamental knowledge and practical abilities in MATLAB required to effectively utilize this tool in technical numerical computations and visualisation in other courses. Syntax and interactive computations, programming in MATLAB using scripts and functions, rudimentary algebra and analysis. One- and two-dimensional graphical presentations. Examples on engineering applications.
8	Outline syllabus	**Introduction to MATLAB**
	Unit 1	**Introduction**
	A	Vector and matrix generation, Subscripting and the colon notation.
	B	Matrix and array operations and their manipulations,
	C	Introduction to some inbuilt functions.
	Unit 2	**Relational and Logical Operators**
	A	Flow control using various statement and loops including If-End statement, If-Else –End statement

	B	Nested If-Else-End Statement,		
	C	For – End and While-End loops with break commands.		
	Unit 3	**m-files**		
	A	Scripts and functions		
	B	concept of local and global variable		
	C	Few examples of in-built functions, editing, saving m-files.		
	Unit 4	**Two dimensional Graphics**		
	A	Basic Plots, Change in axes and annotation in a figure		
	B	multiple plots in a figure		
	C	saving and printing figures		
	Unit 5	**Applications of MATLAB**		
	A	Solving a linear system of equations,		
	B	Curve fitting with polynomials using inbuilt function such as polyfit, solving equations in one variable,		
	C	Solving ordinary differential equations using inbuilt functions		
	Mode of examination	Theory		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book	An introduction to MATLAB : Amos Gilat		
	Other References	1. Applied Numerical Methods with Matlab for engineering and Scientists by stevenchapra, Mcgraw Hill. 2. Getting started with Matlab: RudraPratap		

2.1 Template A1:Inorganic Chemistry-II (MCH135)

School: SBSR		**Batch : 2020-22**
Program: M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester: II**
	Course Code	**MCH135**
	Course Title	**Inorganic Chemistry II**
	Credits	**4**
	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
	Course Objective	1.To introduce the basics concept of molecular symmetry and group theory 2.To demonstrate the various application of group theory in spectroscopy 3.To provide an introduction to basic concepts of organometallic chemistry 4.To explain to the student the various application of organometallic chemistry in industry 5.To provide information various industrially important organometallic compounds. 6.To provide structure, bonding and reactivity of transition metal carbonyls, nitrosyls and phosphin complexes.
	Course Outcomes	CO1:Understand the various basics concept of molecular symmetry and group theory. CO2:Apply their knowledge of group theory to understand the principles of spectroscopy. CO3:Know the basic concepts of organometallic chemistry and its application in industry. CO4: Explain the structure and reactivity of transition metal alkyl, aryl, alkene, alkynes, allyls, dienyl and arene and carbine complexes. CO5: Gain insight about transition metal carbonyls, nitrosyls and phosphin complexes. CO6: Gain knowledge about advanced topics like organometallic chemistry and group theory.
	Course Description	The course includes the basic concept of group theory and its application in chemistry; as well as organometallic chemistry of transition metals.
	Outline syllabus	
	Unit 1	**Molecular symmetry**
	A	Introduction, Meaning and examples of different symmetry elements and generated operations; and general rules, Derivation of matrices for rotation; reflection; rotation ; reflection and inversion operations;

	B	Symmetry operations of all the molecular point groups (C_n, D_n, C_{nh} , $D_{nb}C_{nv}$, D_{nd}, S_n, T, T_d, T_h, O, O_h, I and I_h); Determination of the classes of operations by similarity transform method (only C_{2v}, C_{2h}, C_{3v}, S_4) and general rules
	C	Defining properties of 'group'; Types of groups (Isomorphic, Cyclic and Abelion); Subgroups; reducible and irreducible representations;
	Unit 2	**Application of Group Theory**
	A	Great Orthogonality Theorem, construction of character table for C_{2v} and C_{3v} point group
	B	Optical activity and dipole moment
	C	Application of group theory to electronic and vibrational spectroscopy
	Unit 3	**Organometallic Chemistry-I**
	A	General Characteristics oforganometallic compounds, Ligand hapticity, electron count for different types of organometallic compounds, 16 and 18 electron rule and exceptions, Fluxionality in organometallic complexes. Stereochemical non-rigidity in organometallic compounds.
	B	Synthesis, structure and bonding of organolithium and organomagnesium compounds
	C	Organometallic reagents in organic synthesis and in homogeneous catalytic reactions (Hydrogenation, hydroformylation, isomerisation, polymerisation and metathesis).
	Unit 4	**Organometallic Chemistry-II**
	A	General synthetic routes, nature of bond and structural characteristics of alkyl, aryl, alkene, alkynes, allyls, dienyl and arene complexes of transition metals.
	B	Structure and bonding of metallocenes.
	C	Synthesis, structure and reactivity of metal carbene and carbynes
	Unit 5	**Organometallic Chemistry-III**
	A	Ligand behavior of CO, General methods of preparation, structures, bonding, and vibrational spectra of metal (Fe, Ru, Os, Cr, Ni) carbonyls.
	B	Ligand behavior of NO (NO^+, NO^- and bridging NO), preparation, structures, bonding and important reactions of nitrosyls of Cr, Fe and Ru
	C	Preparation, structure, bonding and reactivity of metal phosphines. Comparison of phosphine and carbonyl ligands in terms of bonding.
	Mode of examination	Theory/Jury/Practical/Viva
	Weightage Distribution	CA: 30% | MTE: 20% | ETE: 50%
	Text book/s*	1. Inorganic Chemistry, J.E. Huhey, Harper & Row. 2.Organometallic Chemistry, R.C.Mehrotra and A.Singh, New Age International.
	Other References	1. Advanced Inorganic Chemistry, F.A. Cotton and Wilkinson, John Wiley 2. Introduction to Ligand fields, B.N. Figgis, Wiley, New York. 3. The Organometallic Chemistry of the Transit ion Metals, R.H. Crabtree, John Wiley. 4. Transition metal chemistry, Fundamental concept and applications, A.Yamamoto, John Wiley, 1986.

2.1 Template A1: Organic Chemistry-II (MCH136)

School: SBSR		**Batch : 2020-22**
Program:M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester:II**
1	**Course No.**	**MCH136**
2	**Course Title**	**Organic Chemistry II**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objective	1. To conceptualize the critical C-C bond forming reactions and in organic synthesis and molecular rearrangements using enolates/ enamines/ metal catalyst or orgaganometallic compounds.. 2. To develop the critical thinking to analyze the conditions required for C=C bond formation 3. To discuss the mechanism of various famous name reactions. 4. To elaborate the process of oxidation and reduction in organic reactions by giving the example of suitable name reactions and develop understanding of the functional mode of different oxidation reduction reagents. 5. To recognize the factors that drives a reactant to undergo rearrangement reaction and understand the different name reactions involving rearrangement.
6	Course Outcomes	The students will be able to- 1. compile the different ways to form C-C bond and associated name reactions. 2. formulate his/her own reasoned opinions in the mechanistic side of C=C bond forming organic reactions 3. enlist a number of oxidizing reagents and analyze the change in oxidation state during the oxidation reaction. 4. understand the functional mode of various reducing reagents. 5. various name reactions and popular rearrangement reactions. 6. develop critical thinking and deep understanding of mechanistic pathways of vast variety of reactions involving new formation, reduction, oxidation and rearrangement reactions.
7	Course Description	This course utilizes the basics developed in organic chemistry to understand the mechanism and in-depth understanding of bond forming (C-C or C=C), Redox, Rearrangement and important name reactions.

2.1 Template A1: Physical Chemistry-II (MCH137)

School: SBSR		**Batch : 2020-22**
Program:M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester:II**
1	**Course Code**	**MCH137**
2	**Course Title**	**Physical Chemistry II**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objective	1. To familiarise students with theoretical and mathematical aspects of quantised energy levels of particle in box, 2. To introduce the theoretical concept of Hydrogen atom and hydrogen molecule and hydrogen molecule ion. 3. To infer the concept of Charge on colloids, electro kinetic phenomenon's and different theories on colloids 4. To prioritise the surface phenomenon's and different equations and theories to explain them. 5. To describe equilibrium processes of one and more than one component systems such as congruent, Peritectic and Monotectic Systems.
6	Course Outcomes	CO1:The concepts of quantum mechanics and its mathematical interpretation for atoms and molecules possessing single electron. CO2:The results and their analysis obtained on the basis of MOT and VBT for hydrogen atom, molecule and ion. CO3:The nomenclature of particles on the basis of particle size and different theories and results related to stability of colloids. CO4:The concept of surface tension, micellization and solubilisation. CO5: The concept of existence of different phases with change in different variables by visualizing the phase diagrams CO6: The concept of quantum mechanics, their application to MOT and VBT, how to draw phase diagrams and importance of colloids and surface chemistry in daily life, their concepts, phenomenon and mathematical equations.
7	Course Description	Concept of Quantum mechanics and its applications in MOT and VBT were shared with students. Theories of colloids and concepts of surface chemistry were discussed. The phase diagram of different component systems were discussed and explained how to plot them.
8	Outline syllabus	
	Unit 1	**Quantum Mechanics**
	A	Matter waves, The Uncertainty principle, The wave nature of the electron, Interpretation of wave function, Normalized and orthogonal wave functions, Linear and Hermitian operators, Commutation of operators, Eigen value and Eigen function

	B	The wave equation, Particle in one dimensional box, particle in three dimensional box, particle in a ring, Degeneracy.Angular momentum operator, Ladder operator,
	C	Hydrogen atom: Schrodinger wave equation, Transformation of coordinates, separation of variable in polar spherical coordinates and its solution, principal, azimuthal and magnetic quantum numbers and their magnitude, probability distribution function, radial distribution function and shape of atomic orbital's (s,p & d), Virial theorem.
	Unit 2	**Chemical Bonding**
	A	Born Oppenheimer Approximation, The ionic bond, The variation method, Ground state energy of the hydrogen atom,
	B	Huckel molecular orbital theory of conjugated systems, delocalisation energy and Secular equations , Molecular orbital theory – Hydrogen molecule ion,
	C	Valence bond theory- Hydrogen molecule, Simple homo and hetero nuclear diatomic molecules, Electronic spectra, effect of substituent on spectra.
	Unit 3	**Colloids**
	A	Introduction, Origin of the charges, electro-kinetic phenomena, electrophoresis, electro osmosis, sedimentation and streaming potential.
	B	The concept of electrical double layer and various models to explain its structure and properties,
	C	DLVO theory and stability of colloids. Smoluchowski theory of kinetics of coagulation and distribution of colloids aggregates. Organic and inorganic gels and clay colloids.
	Unit 4	**Surface Chemistry and Micelles**
	A	**Surface** tension and surface free energy; Pressure across an interface: Laplace equation, Kelvin equation; Wetting: Young-Dupre equation;
	B	Adsorption in liquid systems: Gibbs adsorption isotherm; Adsorption on solids: Langmuir isotherm, BET isotherm, transition state theory of surface reactions: rates of chemisorption and desorption.
	C	Micelles-Surface active agents and their classification, micellization, hydrophobic interaction, critical micellar concentration (cmc), factors affecting cmc of surfactants, counter ion binding to micelles, thermodynamics of micellization-phase separation and mass action models, solubilization, micro emulsions, reverse micelles.
	Unit 5	**Phase Equilibria**
	A	Statement and meaning of the terms in Gibbs phase rule; Thermodynamic derivation of Gibb's phase rule, phase equilibria of water, Hellium and carbon systems;
	B	Two component solid-liquid equilibria (example of Cu-Ni alloy,Bi - Cd system and $CuSO_4 – H_2O$ System): simple eutectic; congruent melting type; peritectic type and monotectic type phase diagrams,
	C	concept of Phase equilibria of three component systems - CaO-Al_2O_3-SiO_2 and Acetic acid-water-Butanol system, Phase-Transformations in Solids: Thermodynamic Classifications of Phase Transitions

	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1.Physical Chemistry, P. W. Atkins, Oxford University Press, New York. 2.Physical Chemistry, I.N. Levine, Tata McGraw Hill Pub. Co. Ltd., New Delhi. 3. Physical Chemistry of Surfaces by A. W. Adamson, John Wiley and Sons.		
	Other References	1.Theoretical Inorganic Chemistry by M.C.Day and J.Selbin 2. Applied Colloid and Surface Chemistry by R. M. Pashley and M. E. Karaman, Wiley Publications. 4.Comprehensive Physical Chemistry by N.B.Singh, N.S.Gajbhiye and S.S.Das , New Age publishers, New Delhi 5.Physical Chemistry by D.A.McQuarrie and J.D.Simon		

2.1 Template A1: Analytical Chemistry-II (MCH138)

School: SBSR		**Batch : 2020-22**
Program: M.Sc.		**Current Academic Year: 2020-22**
Branch: Chemistry		**Semester: II**
1	**Course Code**	**MCH138**
2	**Course Title**	**Analytical Chemistry II**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objective	1. Understand the theories and principles of qualitative and quantitative analysis through optical and spectroscopic technique. 2.Analyse the textural information of bulk materials and particle dimension. 3.Carry out qualitative and quantitative analysis employing descriptive knowledge of electrochemistry and electrochemical titration. 4.Separate and estimate macromolecule (proteins, enzymes, blood and natural products) electroanalytically. 5.Effectively use various sensors for estimation and gain idea about developing technologically potent sensor materials. 6. To learn the advance spectroscopic and microscopic methods for the analysis of molecular materials.
6	Course Outcomes	CO1: Understand various optical and spectroscopic methods for qualitative and quantitative analysis of metals and non metal to trace level. CO2: Evaluate the properties of materials such as porosity, density and microstructure of materials. CO3: Develop new synthetic routes involving electrochemical redox process. CO4: Understand principles of Cyclic Voltammetry and Electrophoresis. CO5: Develop quick, sensitive and selective sensory materials for qualitative and quantitative estimation of analyte. CO6: Investigate the molecular materials using advanced spectroscopic and microscopic techniques.
7	Course Description	Analytical chemistry II emphasizes on various parts of analytical methods as - Atomic Spectroscopy comprises of AAS, AES and ICPMS, Electron Microscopic techniques comprises of SEM, TEM and FESEM, Polarography and amperometry, Cyclic voltammetry and electrophoresis Chemical sensors
8	Outline syllabus	
	Unit 1	**Atomic Spectroscopy**
	A	Theory, sources, burners, atomic emission spectra, atomic absorption spectra, effect of temperature on emission and absorption,

		Instrumentation for AES and AAS, standard addition and internal standard method of analysis		
	B	Comparison of atomic absorption and emission methods, Applications of AAS and AES Features of atomic mass spectroscopy, Atomic weight in mass spectroscopy, mass to charge ratio		
	C	Types of atomic mass spectroscopy, quadruple mass analyzer, time of flight mass analyzer, Inductively coupled mass spectroscopy (ICPMS), Instrumentation for ICPMS, Applications of ICPMS		
	Unit 2	**Electron Microscopic Techniques**		
	A	Basic principle, instrumentation and application of Transmission Electron Microscope (TEM) and HRTEM		
	B	Basic principle, instrumentation and application of Scanning Electron Microscope (SEM)		
	C	Basic principle, instrumentation and application of FESEM		
	Unit 3	**Electroanalytical Technique I**		
	A	**Polarography** Introduction, Instrumentation, Ilkovic equation and its verification		
	B	Derivation of wave equation, Determination of half wave potential, qualitative and quantitative applications		
	C	**Amperometry:** Basic principles, instrumentation, nature of titration curves and analytical principles		
	Unit 4	**Electroanalytical Technique II**		
	A	**Cyclic Voltammetry** Cell design, instrumentation, current-potential relation for linear sweep voltammetry (LSV), cyclic voltammetry, interpretation of voltammograms.		
	B	**Electrophoresis:** Separation by adsorption-Affinity techniques, affinity elution from ion exchangers and other adsorbents		
	C	Pseudo affinity adsorbents, polyacrylamide get electrophoresis, isoelecrictric focusing, isotachophoresis		
	Unit 5	**Chemical Sensors**		
	A	Principles, types of chemical sensors based on the modes of transductions, Types of chemical sensor based on the chemically sensitive materials		
	B	solid electrolyte, gas, semiconductor, Humidity sensors, Biosensors sensors		
	C	Electrochemical sensors (Potentiometric sensors, Ion-selective electrodes, Membrane electrodes, Amperometric sensors)		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	Principles of Instrumental Analysis, Skkog, Holler, Nieman, (Sixth Ed.)		
	Other References	1) Introduction to Instrumental Analysis by R. D. Broun, Mc Graw Hill (1987)		

		2) Instrumental methods of chemical analysis by H. willard, L.Merrit, J.A. Dean and F.A. settle. Sixth edition CBS (1986) 3) Fundamentals of Analytical Chemistry, 6th edition, D.A. Skoog, D.M. West and F.J. Holler, Saunders college publishing. 4) Principles of Instrumental Analysis, Skkog, Holler, Nieman, (Sixth Ed.) 5) Introduction to instrumental analysis by R. D. Braun, MC. Graw Hill-International edition. 6) Analytical Chemistry, Ed. by Kellner, Mermet, otto, Valcarcel, Widmer, Second Ed. Wiley –VCH 7) Electron microscopy in the study of material, P. J Grundy and G. A Jones, Edward Arnold.

2.1 Template A1: Renewable Energy Resources (MPH115)

School: SBSR		**Batch : 2020-2022**
Program: M.Sc		**Current Academic Year: 2020-22**
Branch: Chemistry		**Semester: II**
1	**Course Code**	**MPH115**
2	**Course Title**	**Renewable Energy Sources**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objective	1. 1. To know the importance of Physics and Materials Science. 2. To utilize the various synthesis procedure to develop materials. 3. To explain the practical application of materials in various area
6	Course Outcomes	CO1: Learn the basics of Materials/Technology CO2: Understand the correlation between Applied science and Technology CO3: Apply the concept of materials and technology at certain levels. CO4: Develop devices using materials. CO5: Create the path to handle materials. CO6: Expertise in various tools will make a bridge between industry and students and find out the platform for employment in high tech industries
7	Course Description	This course is based on renewable energy that is collected from renewable resources, which are naturally replenished on a human timescale, such as sunlight, wind, rain, tides, waves, and geothermal heat. Renewable energy often provides energy in four important areas: electricity generation, air and water heating/cooling, transportation, and rural (off-grid) energy services
8	Outline syllabus	
	Unit 1	**Natural and Renewable Energy Resources**
	A	Natural resources and associated problems, Forest, Water, Mineral, Food, Energy and Land resources
	B	Use and over-exploitation, Concept of an ecosystem, Environmental Pollution, Nuclear hazards
	C	Renewable Energy sources: Definition and types of renewable sources, Wind, Ocean, Geothermal, Biomass, Hydro as renewable energy resources
	Unit 2	**Solar Energy: Fundamental and Material Aspects**
	A	Fundamentals of photovoltaic Energy Conversion Physics and Material Properties, Types of solar energy conversion
	B	solar thermal: basics and design of water heaters, solar ponds, Basic to Photovoltaic Energy Conversion: Optical properties of Solids
	C	Direct and indirect transition semiconductors, interrelationship between absorption coefficients and band gap recombination of carriers.

	Unit 3	**Solar Energy: Different Types of Solar Cells**		
	A	Types of Solar Cells, p-n junction solar cell, Transport Equation, Current Density, Open circuit voltage and short circuit current		
	B	Brief description of single crystal silicon and organic and Polymer Solar Cells, Elementary Ideas of Advanced Solar Cells e.g. Tandem Solar cells, Solid Liquid Junction Solar Cells		
	C	Nature of Semiconductor, Principles of Photo-electrochemical Solar Cells.		
	Unit 4	**Hydrogen Energy: Fundamentals, Production and Storage**		
	A	Hydrogen as a source of energy, Solar Hydrogen through Photoelectrolysis, Physics of material characteristics for production of Solar Hydrogen		
	B	Brief discussion of various storage processes, special features of solid hydrogen storage materials		
	C	Structural and electronic characteristics of storage material, New Storage Modes.		
	Unit 5	**Hydrogen Energy: Safety and Utilization**		
	A	Various factors relevant to safety, use of Hydrogen as Fuel, Use in Vehicular transport, Hydrogen for Electricity Generation		
	B	Fuel Cells, Various type of Fuel Cells, Applications of Fuel Cell		
	C	Elementary concepts of other Hydrogen- Based devices such as Hydride Batteries		
	Mode of examination	Theory		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1.Fundamentals of Solar Cells Photovoltaic Solar Energy :Fahrenbruch&Bube		
	Other References	1.Solar Cell Devices-Physics :Fonash 2. Phoptoelectrochemical Solar Cells: Chandra 3. Hydrogen as an Energy Carrier Technologies Systems Economy : Winter &Nitch (Eds.) 4. Hydrogen as a Future EngeryCarrier : Andreas Zuttel, Andreas Borgschulte and Louis Schlapbach		

2.1 Template A1: Molecular Spectroscopy (MCH231)

School: SBSR		**Batch : 2020-22**
Program:M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester:III**
1	**Course No.**	**MCH231**
2	**Course Title**	**Molecular Spectroscopy**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objective	1.To know the principle and applications of molecular spectroscopy. 2.To understand the theories of UV, FT-IR, Raman, NMR, and Mass spectroscopic techniques. 3.Analyze and identify simple organic molecules by using UV, IR, Mass, ^{1}H NMR and ^{13}C NMR data. 4.To evaluate the application of NMR and Mass spectroscopic techniques to different molecules. 5.To know the principle and instrumentation of different spectrophotometric techniques. 6.To impart the knowledge of electronic, rotation, vibration. NMR, FTIR, ESR, spectroscopy and their applications..
6	Course Outcomes	CO1:Explain the general principles and theory of spectroscopy, distinguish the specialities and applications of various types of spectroscopic methods. CO2:Describe the concept and instrumentation of atomic uv-visible absorption, infrared NMR and Mass spectrometers. CO3:Apply Woodward Fieser Rules. CO4:Understand first and second order ^{1}HNMR spectra. CO5:Solve analytical science problems involving uv-visible absorption, infrared ^{1}H, ^{13}C and mass techniques. CO6:Predict UV, IR, Proton chemical shift, spin-spin coupling, coupling constants and apply ^{13}C resonance spectroscopy and mass spectroscopy to chemical structures.
7	Course Description	The course is framed to give fundamental concepts of UV-Visible, IR, ^{1}HNMR, ^{13}CNMR and Mass spectroscopy. Applications of these spectroscopic techniques to organic/inorganic systems will be discussed.
8	Outline syllabus	
	Unit 1	**UV-Visible Spectroscopy**
	A	Lamberts–Beers law, Electronic spectra, Frank-Condon Principle, predissociation spectra, Fortrat diagram,
	B	conjugated polyene and enone systems, and different types of charge transfer transitions and their basis
	C	Charge transfer spectra in organic and inorganic systems
	Unit 2	**Infrared Spectroscopy**

	A	Basic principle and sample handling. Modes of stretching and bending, bond properties and absorption trends,		
	B	Survey of vibrational frequencies of alkanes, alkenes, alkynes, aromatic compounds, alcohols, ethers, phenols and amines. ketones, aldehydes, esters, amides, acids, anhydrides, lactones, lactams and conjugated carbonyl compounds.		
	C	Effect of hydrogen bonding and solvent effect on vibrational frequencies, overtones, combination bands and Fermi resonance.		
	Unit 3	**Nuclear Magnetic Resonance Spectroscopy-I**		
	A	^{1}H NMR - Effect of magnetic field strength on sensitivity and resolution, chemical shift δ, inductive and anisotropic effects on δ, chemical structure correlations of δ, chemical and magnetic equivalence of spins, spin-spin coupling, structural correlation to coupling constant J		
	B	first order and second order spectra, examples of AB, AX, ABX, AMX and AA′BB′ systems, simplification of second order spectrum, selective decoupling, double resonance; classification of splitting pattern; spin; de coupling; chemical exchange; effect of deuteration		
	C	Structural elucidation of organic compounds using ^{1}H NMR technique		
	Unit 4	**Nuclear Magnetic Resonance Spectroscopy-II**		
	A	^{13}C NMR- Introduction, interpretation of ^{13}C NMR spectra, Chemical shifts and its calculation,		
	B	proton coupled and decoupled spin-spin splitting; Application of DEPT technique to the analysis of CH multiplicities in ^{13}C NMR spectroscopy. Correlation spectroscopy - Illustration of practical applications of ^{1}H-^{1}H COSY, ^{1}H-^{13}C COSY.		
	C	Nuclear overhauser enhancement (NOE).Basic concept of Heternonuclear (F, P, Si) NMR.		
	Unit 5	**Mass Spectrometry**		
	A	Measurement technique (El; FAB); Resolution; exact masses of nucleides; molecular ions; isotope ions; fragment ions of odd and even electron types; rearrangement ions		
	B	factors affecting cleavage patterns; simple cleavage; cleavage at a hetero atom; multi centre fragmentation		
	C	Structure elucidation of organic compounds employing mass spectroscopy; Special methods of GCMS; High resolution MS.		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text Book/s*	1.Spectroscopy of Organic Compounds – P.S.Kalsi, 6th edition, 2004. 2.Molecular Spectroscopy – Banwell, 5th Edition, 2013		
	Other References	1.Applications of Absorption Spectroscopy of Organic Compounds – Dyer, 1st Edition, 2009. 2.Spectroscopic Methods in Organic Chemistry by D.H. Williams and I. Fleming, 4th edition, Tata McGraw-Hill Publishing company Ltd., New		

		Delhi. 3.Spectrometric Identification of Organic Compounds- R. M. Silverstein, F. X. Webster, D. Kiemle, 7th Edition, 2005. 4.Physical Methods in Inorganic Chemistry by R. S. Drago, Affiliated East-West Press, 1st Edition. 5.Spectroscopic identification of organic compounds by Kiemle Webster Silverstein, 7 2nd Edition, 2005

2.1 Template A1: Inorganic Chemistry-III (MCH232)

School: SBSR		**Batch 2020-22**
Program: M.Sc.		**Current Academic Year : 2020-22**
Branch: Chemistry		**Semester III**
1	**Course No**	**MCH232**
2	**Course Title**	**Inorganic Chemistry III**
3	Credits	4
4	Contact hours(L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objectives	1.To explain the reaction mechanism of an inorganic reaction. 2.To discuss factors affecting stability of complexes. 3.To explain the route of addition of molecules in a reaction. 4.To have an overview of chemistry of CO complexes. 5.To explain the concept of stereoisomerism in inorganic complexes. 6.To demonstrate mechanisms of substitution reaction and compare it with associative reaction.
6	Course Outcome	CO1: Explain the trends of rate constants and its determination with different methods. CO2: Provide explanation for substitution in octahedral and square planar complexes. CO3: Explain ligand replacement reactions under different conditions. CO4: Distinguish between oxidative addition and reductive elimination mechanisms. CO5: Analyze the chemistry of carbonyl compounds and metal hydrides. CO6: Gain knowledge about various aspects of inorganic reaction mechanism
7	Course Description	The course gives a detailed view of reaction mechanism, electron transfer mechanisms, oxidative addition and insertion reactions of transition metal complexes.
8	Outline syllabus	
	Unit 1	**Reaction Mechanism of Transition metal complexes-I**
	A	Rate Law, Steady state, Activated complex theory. Stepwise and overall formation constants, their interaction
	B	determination of formation constant by pH-meter, Job's method and spectrophotometery. Trends in stepwise constants
	C	factors affecting the stability of metal complexes with reference to the nature of metal ion and ligand. Chelate effect and its thermodynamic origin
	Unit 2	**Reaction Mechanism of Transition metal complexes-II**
	A	Inert and labile complexes, mechanisms of substitution reactions (dissociative, associative interchange mechanism), the conjugate mechanism,
	B	direct and indirect evidence in favour of conjugate mechanism,

		substitution in cis and trans complexes, isomerism of chelate rings, *trans* effects, explanation for *trans* effect		
	C	Ligand replacement reactions of square planar and octahedral complexes: their factors and mechanism of substitution, Anation reactions.		
	Unit 3	**Electron Transfer Mechanisms**		
	A	Inner sphere and outer sphere reactions and their mechanisms		
	B	Racemization and Isomerization, Effect of ligand field on reaction rates		
	C	Mixed valence complexes, Marcus-Husch theory, Thermal and optical electron transfer reactions.		
	Unit 4	**Oxidative-Addition and Migration (Insertion Reactions)**		
	A	Introduction: Acid base behaviour of metal atoms in complexes, Protonation and Lewis Base behaviour, acceptor properties of Lewis acidity of complexes		
	B	oxidative addition and reductive elimination, addition of specific molecules, Hydrogen addition, HX additions, Organic halides addition of some other molecules productive elimination, migration (Insertion) reaction		
	C	promotion of alkyl migration, insertion of CO into M-H bonds, other aspects of CO insertion reactions, Insertion of alkenes and C-C unsaturated compounds, Cleavage of C-H bonds; alkane activation, Cyclometallation reactions.		
	Unit 5	**Metal Hydride Complexes**		
	A	Synthesis, structure and reactions of hydrido complexes, characterization of complexes, molecular hydrogen compounds-synthesis and reactions		
	B	Mononuclear polyhydrides, homoleptic polyhydride anions; carbonyl hydrides and anion		
	C	MH interactions; synthetic applications of metal hydrides		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1.J.E.Huheey. Inorganic Chemistry: Principles of Structure and Reactivity. Harper Inter science.		
	Other References	1.William L. Jolly, Modern Inorganic Chemistry, 2nd Edn, Tata McGraw Hill. 2.E. A. V. Ebsworth, D. W. H. Rankin and S. J. Cradock. Structural methods in Inorganic Chemistry, Blackwell Scientific Oxford. 3.I. P. Atkins, T. Overton, J. Rourke, M. Weller, F. Armstrong, Shriver and Atkins. Inorganic Chemistry, Oxford University Press. 4.T. Moeller. Inorganic Chemistry: A Modern approach, John Wiley. 5.F. Basalo and R.G.Pearson, Mechanism of Inorganic reactions,2nd Edn ,Wiley Eastern Ltd., New Delhi		

2.1 Template A1:Physical Chemistry-III (MCH233)

School: SBSR		**Batch 2020-22**
Program: M.Sc.		**Current Academic Year : 2020-22**
Branch : Chemistry		**Semester III**
1	**Course Code**	**MCH233**
2	**Course Title**	**Physical Chemistry III**
3	Credits	4
4	Contact hours	4-0-0
	Course Status	Compulsory
5	Course Objectives	The main objectives of this program is to: 1: To provide deep knowledge on advanced quantum chemistry. 2: To provide a thorough proficiency in approximate methods in quantum chemistry. 3: To enable students to interpret many electron systems quantum mechanically. 4: To impart knowledge on kinetics of complex reactions. 5: To make the student understand the kinetics of reaction in solution . 6: Apply the knowledge about quantum chemistry and kinetics to solve real life problems .
6	Course Outcome	After successful completion of the course, the students will be able to: CO1: understand different polynomials and their application. CO2. apply the knowledge of time dependent perturbation theory and variational method for quantum mechanical problems. CO3 apply the quantum chemistry knowledge to analyse the behaviour of multi electron systems. CO4. explain the kinetics of various types of complex reactions CO5. Apply the knowledge of kinetics of reactions in solution to solve kinetics problems. CO6.Apply knowledge quantum chemistry to solve real life problems and kinetics to understand mechanism of reactions.
7	Course Description	
8	Outline Syllabus	
	Unit 1	Advanced Quantum chemistry: Prerequisite
	A	Legendre, associated Legendre polynomials; Hermite polynomials; Lagurre and associated Lagurre polynomials; polynomials as orthonormal functions, their properties; step-up and step-down operators, application to single electron and multi-electron atom,
	B	eigen-ket-ladder and formulation of spherical harmonics from angular momentum rules, finite rotation operation vs. angular momentum operators, spin angular momentum, Pauli spin matrices — spin eigenfunctions and their properties.

	C	coupling of angular momentum for many electron system, spin-orbit coupling, Molecular term symbols. Quantum tunnel effect. Fermi and Bose gases.
	Unit 2	Approximate methods
	A	Time dependent perturbation theory, semi classical treatment of radiation-matter interaction, transition probability and rates, Einstein's A and B coefficients, selection rules; Oscillator strength,
	B	Variation theorem and variational methods: principles of linear and non-linear variation methods,
	C	stationary perturbation theory for non-degenerate and degenerate states - applications to rotator, Stark effect.
	Unit 3	Many electron systems
	A	Antisymmetry of many electron wave function, spin and spatial orbitals, Slater determinant; closed-shell and open-shell electron configurations; multi-electron pure-spin state wave functions - examples with 2- and 3-electron systems,
	B	formulation of a multi-electron closed-shell electron configuration energy, introduction of core, Coulomb and exchange integrals with their properties - example of He atom, independent particle model, multi-electron atomic Hartree Hamiltonian and related SCF equations solution,
	C	Roothaan-Hartree-Fock method vertical ionization potential and Koopman's theorem; Problems with open-shell systems. Restricted and unrestricted HF methods (elementary idea). discussion of electron correlation.
	Unit 4	Kinetics of complex reactions
	A	Application of statistical mechanics to transition state theory, comparison of transition state theory with experimental results, Kinetics of complex reactions (reversible, simultaneous and consecutive),
	B	chain reactions; branched and non-branched kinetic rate equations, population explosion, upper and lower ignition/explosion limits; thermal ignition and ignition temperature; chemical oscillation: conditions for oscillation, chemistry of BZ reaction (Brusselator model); autocatalysis,
	C	Fast reactions, experimental techniques for fast reactions (stopped-flow, temperature- jump and flash photolysis
	Unit 5	Reactions in solution
	A	Reaction between ions, effect of solvent (single & double sphere models), interpretation of frequency factor and entropy of activation, influence of ionic strength, salt effect, reactions involving dipoles,
	B	influence of pressure and volume on reaction rates in solution. Intermolecular potential and centrifugal barrier, impact parameter, collision cross section and rate, energy threshold, opacity function and reaction cross section
	C	Discussion of physicochemical techniques for kinetic study.
	Mode of examination	Theory/Jury/Practical/Viva

	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1. 1. Quantum Chemistry, I.M. Levine, Prentice Hall. 2. 2. Chemical Kinetics, K. J. Laidler, Harper & Row, New York.		
	Other References	1. 1. Quantum Chemistry by D.A.McQuarrie Viva Books 2. *Quantum Chemistry*, H. Eyring, J. Walter and G.E. Kimball, (1944) John Wiley, New York. 3. Foundations of Chemical Kinetics – S.W. Benson		

2.1 Template A1:Organic Chemistry-III (MCH234)

School: SBSR		**Batch 2020-22**
Program: M.Sc.		**Current Academic Year : 2020-22**
Branch : Chemistry		**Semester III**
1	**Course No.**	**MCH234**
2	**Course Title**	**Organic Chemistry III**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objective	1.Oxidation and reduction reagents and their application for functional group conversion in organic synthesis. 2.Explain retro-synthesis of aromatic, alicyclic and aliphatic compounds and synthons. 3.The ability to recognize reagents for functional group transformations. 4.Retrosynthetic simplification of target molecules and to provide forward synthetic proposals. 5.Designing a retrosynthetic approach for the synthesis of a target molecule.
6	Course Outcomes	CO1:Role of various reagents used in organic chemistry. CO2:Have a thorough grounding in protection and deprotection chemistry. CO3:Identify the components of retrosynthesis. CO4:Understand the synthesis and properties of metallocenes, non-benzenoids and polycyclic aromatics. CO5: Design a green synthesis using principles of prevention of waste/by-products/toxic products, atom economy. CO6: Gain in-depth knowledge in synthetic organic chemistry.
7	Course Description	The aim of this organic chemistry course is to provide an in-depth overview of retrosynthetic analysis and the disconnection approach. These are fundamental concepts used by organic chemists in designing the synthesis of target molecules in sectors such as pharmaceuticals, agrochemicals and fine chemicals.
8	Outline syllabus	
	Unit 1	**Reagents in Organic Synthesis**
	A	Use of the following reagents in organic synthesis and functional group transformations; Gilman's reagent, lithium diisopropylamide (LDA), dicyclohexylcarbodiimide(DCC)
	B	1,3-dithiane (reactivity Umpoloung), trimethylsilyl iodide, tri-n-butyltin hydride, DDQ,
	C	Phase transfer catalysts, crown ethers and Merrifield resin, Wilkinson's catalyst, Baker yeast.

	Unit 2	**Protection and Deprotection of Functional Groups**		
	A	Protection and deprotection of hydroxy, carboxyl, carbonyl, carboxy groups		
	B	Protection and deprotection of amino groups and carbon-carbon multiple bonds		
	C	chemo- and regioselective protection and deprotection, illustration of protection and deprotection in multi-step synthesis		
	Unit 3	**Retrosynthetic Analysis**		
	A	Basic principles and terminology of retrosynthesis, guidelines, synthesis of aromatic compounds		
	B	one group and two group C-X disconnections, one group C-C and two group C-C disconnections, amine and alkene synthesis		
	C	important strategies of retrosynthesis, functional group transposition, important functional group interconversions, reversal of polarity (umpolung)		
	Unit 4	**Metallocenes, Non-benzenoid Aromatics and Polycyclic Aromatic compounds**		
	A	General considerations, synthesis and reactions of some representative compounds - tropone, tropolone, azulene,		
	B	General considerations, synthesis and reactions of some representative compounds - ferrocene, fluorene,		
	C	General considerations, synthesis and reactions of some representative compounds - phenanthrene and indene.		
	Unit 5	**Green Chemistry**		
	A	Principles of Green Chemistry, Concept of atom economy, Tools of Green Chemistry: Alternative feedstocks/starting materials, Reagents, Solvents, Product/target molecules, Catalysis and process analytical chemistry.		
	B	Evaluation of chemical product or process for its effect on human health and environment, Evaluation of reaction types and methods to design safer chemicals. Evaluating the effects of Chemistry:		
	C	Toxicity to humans, Toxicity to wildlife, Effects on local environment, Global environmental effects. Planning a green synthesis.		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text Book/s*	1.Organic reactions and their mechanisms, P.S. Kalsi, New Age International. 2.Reagents for Organic Synthesis, L.F. Fieser and M. Fieser. 3.Organic Synthesis: The Disconnection Approach, Stuart Warren, Paul Wyatt. 4. Organic Chemistry, I.L. Finar Volumes I & II.		
	Other references	1. Anastas, P., and Williamson, T. C., Green Chemistry Frontiers in Benign Chemical Synthesis and Processes, Oxford University Press (1999).		

		2.Ahluwalia, V. K., and Kidwai, M., New Trends in Green Chemistry, Anamaya Publishers (2004). 3.Protective Groups in Organic Synthesis, Peter G. M. Wuts, T.W. Greene. 4.Sheldon, R.A., Arends, I., and Hannefed, U., Green Chemistry and Catalysis, Wiley-VCH Verlag GmbH and Co. (2007).

2.1 Template A1:Inorganic Chemistry-IV (MCH235)

School: SBSR		**Batch 2020-22**
Program: M.Sc.		**Current Academic Year : 2020-22**
Branch:Chemistry		**Semester:III**
Course Code		**MCH235**
Course Title		**Inorganic Chemistry IV**
1	Credits	4
2	Contact Hour	4-0-0
	Course Status	Compulsory
5	Course Objective	1.To describe about the structure, properties and uses of inorganic chains. 2. To provide information about inorganic ring compounds. 3. To introduce the basic concepts about cluster structure and their reactivity. 4.To illustrate the basic concepts of inorganic photochemistry. 5.To describe the various photochemistry of various inorganic metal complexes. 6. To know about the application of photochemistry.
6	Course Outcome	CO1: Explain the structure, properties and uses of inorganic cages and chains. CO2: Describe the structure and properties of inorganic rings. CO3: Predict the structure of inorganic clusters using Wade's rule. CO4: Understand photochemical reactions of various coordination compounds. CO5: Apply the knowledge of photochemistry in real life problems. CO6: Gain knowledge about advanced topics like inorganic photochemistry and inorganic clusters
7	Course Description	The course is designed to appraise the chemistry of inorganic chains, cages, rings, clusters. The photochemistry of inorganic compounds is also covered in detail.
8	Outline syllabus	
	Unit 1	**Chains and Cages**
	A	Structural aspects of silicate minerals and silicones, Zeolites-Structure, applications and synthesis, Intercalation Chemistry, One dimensional conductors, (SN)x chains.
	B	Cages: Electron deficient bonding in higher boranes and its derivatives, Types of heteroboranes with special reference to carboranes, structure, bonding and IUPAC nomenclature.
	C	Metallaboranes, metal σ and μ bonded borane/carborane clusters. Resemblance of Metallaboranes with ferrocene and related compounds. Applications of Metallaboranes.
	Unit 2	**Rings and Clusters**
	A	Rings : Synthesis, structure and chemical application of borazine, Phosphazene, phosphazene polymers, Metal-Metal bonds. Concept of quadrupolar bond and its comparison with a C-C bond.

	B	Clusters: Types of metal clusters and multiplicity of M-M bonds. Simple and condensed metal carbonyl clusters-types, calculation of number of M-M bonds using 18/16 electron rule in low and high nuclearity metal clusters, capping rule.		
	C	Application of Wade's rule over metral carbonyl clusters. Metal halide and metal chalcogenide clusters.		
	Unit 3	**Photo Inorganic Chemistry-I**		
	A	Introduction, Absorption, excitation, photochemical laws, quantum yield, electronically excited states, Photochemical laws; Jablonski Diagram		
	B	radiative and non-radiative processes, Franck-Condon principle, photochemical stages-primary and secondary processes, Kasha's rule, Thexi state		
	C	Types of photochemical reactions in transition metal complexes– substitution, decomposition, fragmentation, rearrangement and redox reactions.		
	Unit 4	**Photo Inorganic Chemistry-II**		
	A	Photo substitution reactions of Cr(III)- ammine complexes : Adamson's rules,		
	B	Photochemistry of Co(III) and Rh(III) Ammine Complexes,		
	C	Photochemistry of Ru- Polypyridyl complexes, comparison of Fe(II) and Ru(II) complexes. Ligand photoreactions, photoredox reactions		
	Unit 5	**Applications of Photochemistry**		
	A	Solar Cells, semiconductor supported metal oxide systems, water photolysis.		
	B	Applications of quenching and sensitization techniques in the identification of reactive state in coordination complexes. Photoreactions and solar energy conversions.		
	C	Photochromism, Photocalorimetry, application of photochemistry in lasers.		
	Mode of Examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text Book/s*	1.J.E.Huheey. Inorganic Chemistry: Principles of Structure and Reactivity. Harper Inter science. 2.F. A. Cotton and G. Wilkinson. Advanced Inorganic Chemistry, Wiley InterScience. 3.Concepts of Inorganic Photochemistry, A. W. Adamson and P. D. Fleischauer, Wiley. 4. Advanced Inorganic Chemistry Vol-1 & 2, Gurdeep Raj, Krishna Prakashan.		
	Other References	1.G. L. Miessler, D. A. Tarr, Inorganic Chemistry, 3rd edition, Pearson Education.		

2.1 Template A1: Physical Chemistry-IV (MCH236)

School: SBSR		**Batch 2020-22**
Program: M.Sc.		**Current Academic Year : 2020-22**
Branch : Chemistry		**Semester III**
1	**Course Code**	**MCH236**
1	**Course Title**	**Physical Chemistry IV**
3	**Credits**	**4**
4	**Contact hours**	**4-0-0**
	Course Status	**Compulsory**
5	Course Objectives	The main objectives of this course is to: 1: To provide the details of advanced topics of spectroscopy. 2. To provide the detailed understanding of Rotational spectroscopy. 3. To provide the structure elucidation methods using IR spectroscopy. 4. To provide the detailed knowledge of the electric structure of molecules. 5. To provide the knowledge of the phenomenon associated with photoelectron spectroscopy. 6. To enrich the student level of understanding of molecular spectroscopy.
6	Course Outcome	After successful completion of the course, the students will be able to: CO1: Analyse the essential parameters from absorption and emission spectrum. CO2: Analyse the microwave spectrum of a molecule. CO3: Analyse the IR spectrum and obtain the bond strength parameters. CO4: Analyse the ground and excited state Absorption and emission spectrum of the molecules. CO5: Investigate the photoelectron spectrum of the molecules. CO6: Correctly predict the molecular structure and associated properties using various spectroscopic techniques.
7	Course Description	
8	Outline Syllabus	
	Unit 1	**Principles of Spectroscopy**
	A	Electromagnetic radiation, Born-Oppenheimer approximation, Heisenberg's Uncertainty Principle,

	B	Jablonski Diagram, Fourier Transform, Time dependent perturbation, Einstein coefficients. Lambert-Beer's law, Integrated absorption coefficients, Transition dipole moments and general selection rules based on symmetry ideas,
	C	Transition probability, oscillator strength, the integrated absorption coefficient.
	Unit 2	**Introduction to Rotational Spectroscopy**:
	A	Rotational spectroscopy of diatomic molecules based on rigid rotator approximation, Determination of bond lengths and/or atomic masses from microwave data,
	B	Effect of isotopic substitution, Non-rigid rotator, Classification of polyatomic molecules,
	C	Energy levels and spectra of symmetric top molecules and asymmetric top molecules, First order Stark effect, FC principle.
	Unit 3	**Vibrational Spectroscopy:**
	A	Force constant and amplitudes, zero potential energy, Morse Potential, Normal coordinates analysis of homonuclear and heteronuclear diatomic molecules, Extension to polyatomic linear molecules,
	B	Derivation of selection rules for diatomic molecules based on Harmonic oscillator approximation. Anharmonic oscillator, Overtones and combination bands, Dissociation energies from Vibrational data, Vibration-rotation spectra, P, Q and R branches, Breakdown of the Born-Oppenheimer approximation.
	C	**Raman Spectroscopy:** Classical and quantum theories, Stokes and anti-Stokes lines, Polarizability ellipsoids, Rotational and Vibrational Raman spectroscopy, pure rotational Raman Spectrum of a linear molecules. Selection rules-Mutually Exclusion Principle, Polarization of Raman lines.
	Unit 4	**UV-Visible Absorption and Emission Spectroscopy:**
	A	Basic principle, Instrumentation and application of absorption and emission spectroscopy, Electronic spectra, Frank-Condon Principle, predissociation spectra, conjugated polyene and enone systems, different types of charge transfer transitions and their basis, Charge transfer spectra in organic and inorganic systems, solvent effects.
	B	Steady-state fluorescence spectroscopy, Mirror-image symmetry and its violation, Radiative and radiationless deactivation, Fluorescence Quenching (static and Dynamics), Room Temperature Phosphorescence, Time-resolved (Time correlated single photon counting-TCSPC) fluorescence spectroscopy, Fluorescence lifetime measurement,
	C	Introduction to Single molecule fluorescence and fluorescence imaging, Photometric titration, comparison of Luminescence and UV Absorption Methods, Limitation of absorption and emission measurement.
	Unit 5	**Photoelectron Spectroscopy:**
	A	The photoionization processes, Auger and autoionization processes, de-excitaion by fluorescence,
	B	outlines of UPS, XPS and Auger techniques and their applications in

		interpretation of valence and core shell spectra of atoms and molecules,		
	C	Laser Spectroscopy.		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1. Fundamentals of Molecular Spectroscopy, Banwell, 3rd Edition, 2018. 2. Pavia, D. L.; Lampmann, G. M.; Kriz, G. S.; Vyvyan, J. R. Introduction to Spectroscopy Cengage Learning (2014). 3. Barrow, G. M. Introduction to Molecular Spectroscopy McGraw-Hill (1962). 4. Hollas. J. M. Modern Spectroscopy 4th Ed., John Wiley & Sons (2004). 5. Chang, R. Basic Principles of Spectroscopy McGraw-Hill, New York, N.Y. (1970).		
	Other References			

2.1 Template A1: Organic Chemistry-IV (MCH237)

School: SBSR		**Batch : 2020-22**
Program:M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester : III**
1	**Course No.**	**MCH237**
2	**Course Title**	**ORGANIC CHEMISTRY IV**
3	**Credits**	4
4	**Contact Hours (L-T-P)**	4-0-0
	Course status	Compulsory
5	**Course Objective**	1.Define the photochemistry and distinguish absorption and emission process 2.Describe the nature of light. 3.Distinguish between electric and magnetic fields, describe the action of light with matter. 4.Compare between spontaneous and simulated emission. h) Describe the electronic transition i) State photochemistry laws 5.Understanding of some important aspects of pericyclic reactions, to learn the orbital interactions (Woodward Hoffmann rules) in concerted reactions 6.Apply concerted and stepwise reactions in organic synthesis
6	**Course Outcomes**	CO1:Define types of photochemical reactions, list the factors determining reactivity, describe Franck Condon Principle. CO2:Compare between Norish type I and Norish type II, distinguish inter & intra molecular cyclo addition, describe photodissociation reaction. CO3:Learn photorearrangement reactions and compare between types of singlet oxygen reactions. CO4: Know what are pericyclic reactions, learn about classification of pericyclic reactions, identify electrocyclic reaction and evaluate application of Woodward-Hoffmann rules to pericyclic reactions. CO5: identify various theories/rules governing electrocyclic reaction, cycloaddition and sigmatropic shifts and analyze which type of pericyclic mechanism is operative in a reaction. CO6:Understand the concepts involved in organic photochemical reactions, their mechanisms and applications in organic synthesis.
7	**Course Description**	The course is framed to make students familiar with the concepts and applications in two important topics in advanced organic chemistry, namely concerted organic reactions and organic photochemistry. Different methods of analysis of pericyclic reactions to arrive at the Woodword-Hoffmann rules are discussed. This course will uncover all the major topics in pericyclic reactions and organic photochemistry.
8	**Outline syllabus**	
	Unit 1	**Photochemistry Part I**
	A	Introduction, Primary photophysical process of atoms and diatomic

		molecules, spectroscopic notations, Frank condon principle and its applications, rates of absorption and emission, quantum efficiency/quantum yield		
	B	quenching of excited states species, radiationless transition and predissociation, energy transfer processes, Wigner's spin rule		
	C	Woodward Hoffman's rule, mechanistic analysis of photochemical reactions by spectroscopic techniques, sources of high energy radiation, chemical dosimetry, comparison between photo and radiation chemistry.		
	Unit 2	**Photochemistry Part II**		
	A	Photochemistry of Olefins- Cis-trans isomerism, cycloaddition, rearrangements. Reaction of conjugated olefins; di-π-methane rearrangements (including oxa- and aza-).		
	B	Photochemistry of Ketones: Excited state of C=O, Norrish type-I and type-II cleavages.		
	C	Paterno-Buchi reaction, α,β-unsaturated ketones, Rearrangement of cyclohexadienones.		
	Unit 3	**Photochemistry Part III**		
	A	Photochemistry of Aromatic compounds - Photorearrangement of benzene and its derivatives, Photo-Fries reactions of anilides, cycloaddition of benzene, Photo-Fries rearrangement		
	B	Barton reaction, Hunsdiecker reaction, Photochemical oxidations and reductions		
	C	Cycloaddition of singlet molecular oxygen, Oxidative coupling of aromatic compounds, photoreduction by hydrogen abstraction		
	Unit 4	**Pericyclic Reactions I**		
	A	Molecular orbital symmetry, Frontier orbitals of ethylene, 1,3-butadiene, 1,3,5-hexatriene and allyl system.		
	B	Classification of pericyclic reactions. Woodward – Hoffmann correlation diagrams. FMO and PMO approach, transition state (ATS) theory, generalized orbital symmetry (GOS) rule.		
	C	Electrocyclic reactions – conrotatory and disrotatory motions, [4n], [4n+2] and allyl systems, torquoselectivity.		
	Unit 5	**Pericyclic Reactions II**		
	A	Cycloadditions – antarafacial and suprafacial additions, 4n and 4n+2 systems. Regio, enantio and Endo selectivities in Diels-Alder reactions.		
	B	Hetero Diels-Alder reaction, 2+2 addition of ketenes, Dipolar cycloadditions, retrocycloadditions.		
	C	Sigmatropic rearrangements - suprafacial and antarafacial shifts of H, sigmatropic shifts involving carbon moieties. [i, j] - sigmatropic rearrangements (including Walk, Claisen, Cope, oxy and aza-Cope rearrangements).		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%

	Text book/s*	1. Reaction Mechanism in Organic Chemistry; S. M. Mukherji and S. P. Singh. 2.Fundamentals of Photochemistry, K. K. Rohatgi-Mukherjee		
	Other References	1. Modern Synthetic reaction by H. O. House, W.A. Benjamin 2. Advanced Organic Chemistry part B, F.A. Carey & R.J. Sundberg, Plenum Press.		

2.1 Template A1: Environmental Chemistry (MCE201)

School: SBSR		**Batch : 2020-22**
Program: M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester: III**
1	**Course Code**	**MCE201**
2	**Course Title**	**Environmental Chemistry**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Elective
5	Course Objective	1.To introduce the basics knowledge of chemistry of environment. 2.To describe the chemistry of hydrosphere. 3.To provide an introduction to chemistry of soil. 4.To explain to the student the causes of industrial pollution. 5.To provide information environmental toxicology. 6.To illustrate the infamous cases of environment related disasters.
6	Course Outcomes	CO1:Understand the chemistry of atmosphere. CO2:Understand the chemistry of hydrosphere. CO3:Explain the chemistry of soil. CO4: Know about adverse effect of industrialization and possible prevention method CO5:Know about environmental toxicology and a few example of environmental disaster. CO6: Gain knowledge about the chemistry of atmosphere, factors affecting it and possible prevention methods
7	Course Description	This course describes the chemistry of earth atmosphere, soil and water bodies. It also describes the adverse effect of industrial pollution and its possible prevention method.
8	Outline syllabus	
	Unit 1	**Earth's Atmosphere**
	A	Introduction, composition of atmosphere, vertical temperature, heat budget of the earth atmospheric system, vertical stability atmosphere
	B	Bio-distribution of elements. Reactions in atmosphere, Stratospheric chemistry. Chemistry of photochemical smog, Precipitation, Acid rain, Production and removal of nitric acid, Sulphuric acid
	C	Atmospheric aerosols-Sources, Concentrations, Control. Chemistry of global climate. Air sampling techniques, Sources, effects and monitoring of air pollutants by Instrumental methods, Control of air pollution
	Unit 2	**Hydrosphere**
	A	Chemical composition of water bodies – lakes, streams, rivers and wetlands etc Hydrological cycle. Aquatic pollution-inorganic, organic, pesticides, agricultural, industrial and sewage, detergents, oil spills and oil pollutants
	B	water quality parameters-dissolved oxygen, biochemical oxygen demand,

		solids, metals, content of chloride, sulphate, phosphate, nitrate and micro-organisms, water quality standards, Analytical methods for measuring BOD, DO, COD, F, OILS, METALS (As, Cd, Cr, Hg, Pb, Se), Residual chloride and chlorine demand, purification and treatment of water.		
	C	Distribution of species in aquatic systems: Single variable diagrams, Two variable diagrams, Method of calculating pE^o		
	Unit 3	**Soils**		
	A	Chemical composition of the soil, micro and macronutrients, the exploitation of the mineral resources and abuse of the earth		
	B	soil pollution due to natural and artificial agencies and its effects, remedial measures to check the pollution. pollution-fertilizers, pesticides, plastics and metals, waste treatment.		
	C	Humic material–Formation, Composition, Structure determination using spectroscopy, Properties. Radioactive pollution, disposal of radioactive waste		
	Unit 4	**Industrial Pollution**		
	A	Cement, sugar, distillery, drug, paper and pulp, thermal power plants, nuclear power plants, metallurgy, polymers, drugs etc		
	B	radionuclide analysis, disposal of wastes and their management. Waste Water, Treatment of Industrial Waste Water,		
	C	Environmental Impact Assessment process in India.		
	Unit 5	**Environmental Toxicology**		
	A	Chemical solutions to environmental problems, biodegradability,		
	B	principles of decomposition, better industrial processes. Bhopal Gas Tragedy, Chernobyl Disaster, Three Mile Island, Sewozo and Minamata disasters**.**		
	C	Occupational safety Hazard Assessment, MSDS		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1.Environmental Chemistry, A.K.Das. 2.Environmental Chemistry, Samir K. Banerji. 3.Environmental Chemistry H. Kaur, 6th Edn, Pragathi Prakashan, Meerut, 2011. 4.Environmental Pollution Analysis, S. M. Khopkar, New Age International (P) Ltd, 1993.		
	Other References	1.Analysis of Industrial Waste Water, K.H.Mancy and W,.J.Weber Jr. Wiley, Interescience New York, 1971. 2.Environmental Chemistry, L.W. Moore and E. A. Moore, McGraw Hill Publication, New York 3.Environmental Chemistry, Colid Baird. W. H. Freemand and Company, 1995.		

2.1 Template A1: Polymer Science and Technology (MCE202)

School: SBSR		**Batch: 2020-22**
Program: M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester:III**
1	**Course Code**	**MCE201**
2	**Course Title**	**Polymer Science and Technology**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Elective
5	Course Objective	1. To impart knowledge on synthesis of polymers using different polymerization methods/techniques and their characterization. 2. To provide basic understanding on the synthesis and characterization of different types of copolymers and preparation of polymer blends and IPNs. 3. To elaborate on the end-uses of polymers as matrix resins for composites, coatings and adhesives. 4. To disseminate information on advanced polymeric systems and speciality polymers. To describe different processing techniques of polymers and to discuss degradation of polymers and the effect of use of polymers on environment.
6	Course Outcomes	CO1:Basic understanding on synthesis of polymers, determination of molecular weight and characterization of polymers using chemical methods and different instruments. CO2:Concept on the factors influencing the copolymerization of monomers and their microstructure, use of block copolymers as thermoplastic elastomers and compatibilizers for polyblends and knowledge on IPNs and Semi-IPNs. CO3:Knowledge on broad spectrum of end-use of polymers as matrix resins for composites, coatings and adhesives and their applications. CO4:Exposure to advanced polymeric systems such as shape memory polymers, self healing polymers, engineering plastics and inorganic polymers. CO5:Understanding of different polymer processing techniques. CO6:Understanding the synthetic pathways and functional polymers along with factors influencing the degradation of polymers and gaining knowledge on the management of plastics and the environmental impact.
7	Course Description	This elective course on Polymer Science and Technology covers the synthesis and characterization of homopolymers and copolymers, thermoplastic elastomers, polymer blends, interpenetrating polymer network (IPN) structures, polymer matrix composites, adhesives and coatings. This course also covers certain advanced/speciality polymer systems such as shape memory polymers, dentrimers, hyperbranched polymers and inorganic polymers. An insight into polymer processing

		techniques, polymer degradation and recycling also forms part of this course.
8	Outline of syllabus	
	Unit 1	**Synthesis and Characterization of Polymers**
	A	Atom Transfer polymerization, Group Transfer Polymerization, Ring Opening Polymerization. Molecular weight: number average, weight average, viscosity average molecular weight, z-average molecular weight, molecular weight distribution.
	B	Measurement of molecular weight and size: Colligative property measurement, Light scattering, ultracentrifuge, viscosity, Gel Permeation Chromatography, Fractionation of polymers by solubility.
	C	Characterization of polymers: chemical methods, spectroscopic methods, X-ray diffraction, microscopy and thermal analysis.
	Unit 2	**Copolymers, Thermoplastic elastomers, polymer blends and IPNs**
	A	Copolymers: Radical copolymerization - monomer reactivity ratios, Q-e factor, Formation of random, alternating and block copolymers in radical copolymerization based on monomer reactivity. Monomer sequencing (diad and triad structures) in copolymers using NMR spectroscopy.
	B	Thermoplastic elastomers: ABA and $(AB)_n$ type block copolymers as thermoplastic elastomers, their microstructure and applications.
	C	Role of block copolymers as compatibilizers Interpene-trating Polymer Networks (IPNs): Semi-IPNs and full IPNs – Synthesis, characterization and applications.
	Unit 3	**Polymer matrix composites (PMCs), Adhesives and Coatings**
	A	Polymer matrix composites: Matrix resins-epoxy resins, phenolic resins and polyimides, Reinforcement-particulate, short fiber, continuous fiber-glass fibre and carbon fibre, characterization techniques and mechanical properties. Polymer Nano Composites, Aerospace and defence applications of PMCs.
	B	Adhesives: Theory of adhesion, an overview of polymers used as adhesives, high temperature adhesives, evaluation of adhesive properties. Applications of adhesives.
	C	Coatings: Water-borne and solvent based coatings, polymers as binders in paints. Self cleaning coatings. Applications of coatings.
	Unit 4	**Advanced Polymers/Speciality Polymers**
	A	Shape Memory Polymers, Self-Healing Polymers, Dentrimers and hyper-branched polymers, Conducting polymers, Liquid Crystalline Polymers.
	B	Engineering thermoplastics: Polyetherimide, Poly-carbonate.
	C	Inorganic polymers: Polyphosphazene, polysilane, polycarbosilane, polysiloxane and polymetallosiloxanes.
	Unit 5	**Polymer Processing, Polymer degradation and the environment**
	A	Basic processing operations: Extrusion, Molding, Coating, Vulcanization and Fiber drawing.
	B	Polymer degradation: Thermal degradation, Oxidative and UV stability,

		Chemical and hydrolytic stability, Effects of radiation.		
	C	Environment: Management of plastics in the environment-recycling, incineration and biodegradation.		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1. Text book of Polymer Science, Third Edition, F.W. Billmeyer, Jr. Wiley-Intersciene, 2003. 2. Polymer Science & Technology, J. R. Fried, Prentice-Hall Inc., USA (Indian Reprint) 2005. 3. Polymers: Chemistry and Physics of Modern Materials, 3rd edition, by J.M.G. Cowie and V. Arrighi, New York, CRC Press, 2008.		
	Other References	1. Macromolecules: An Introduction to Polymer Science, F. A. Bovey and F. H. Winslow, Academic Press, New York, 1979. 2. Inorganic Polymers, 2nd Edtition, J. E. Mark, H. R. Allcock and R. West, Oxford University Press, 2005. 3. Adhesives Technology Handbook, 3rd Edition, Sina Ebnesajjad and Arthur H. Landrock (Imprint: William Andrew) Elsevier, 2014. 4. Processing of Polymer Matrix Composites, P.K. Mallick, CRC Press, 2017. 5. Engineering Thermoplastics: Properties and Applica-tions, Margolis, CRC Press, 1985.		

2.1 Template A1: Inorganic Chemistry-V (MCH238)

School: SBSR		**Batch : 2020-22**
Program: M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester: IV**
1	**Course Code**	**MCH238**
2	**Course Title**	**Inorganic Chemistry V**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objective	1. To describe about basic principles and importance of various metals in natural systems. 2. To describe various ion transport through biological membrane. 3.To explain the importance of Iron and Copper containing metallo-biomolecule. 4.To illustrate the chemistry of bio molecules like DNA and RNA. 5.To describe the bioinorganic chemistry of Molybdenum, Tungsten and Zinc containing Enzymes. 6. To describe the bioinorganic chemistry of Vitamin B_{12}.
6	Course Outcomes	CO1:Explain the transport of ions through membrane CO2: Predict the structure and mechanism of Fe and Cu containing metalloproteins. CO3:Learn about structure and chemistry of DNA and RNA. CO4: Understand the importance of Molybdenum, Tungsten and Zinc containing Enzymes. CO5 : Illustrate biologically important processes like photosynthesis CO6:Understand the role and importance of metal ions in biology.
7	Course Description	This course includes details discussion about various bio molecules and metal containing enzymes with special reference to iron, copper, zinc, tungsten and molybdenum.
8	Outline syllabus	
	Unit 1	**Bioinorganic Chemistry of Metals**
	A	Essential and trace elements in biological systems,
	B	structure and functions of biological membranes; mechanism of ion transport across membranes; sodium pump, role of calcium in muscle contraction, blood clotting mechanism and biological calcification.
	C	Structure and functions of amino acids, proteins, peptides and comparative study of structures and functions of these biomolecules
	Unit 2	**Bioinorganic Chemistry of Iron and Copper**
	A	Iron-sulphur proteins: rubredoxin and ferredoxins;
	B	Metalloporphyrins; Heme proteins: hemoglobin, myoglobin. Cytochrome P-450, Cytochrome c-oxidase and cytochrome c;
	C	Synthetic oxygen carrier and model systems. Thermodynamic and

		kinetics of oxygenation; Non-heme proteins: hemerythrin and hemocyanin.		
	Unit 3	**Bioinorganic Chemistry in Biological Systems**		
	A	Metal complexes of polynucleotides, nucleosides and nucleic acids (DNA and RNA).		
	B	Stability of DNA and melting temperature.		
	C	Role of metal ions in replication and transcription process of nucleic acids. Metal deficiency and disease		
	Unit 4	**Molybdenum, Tungsten and Zinc containing Enzymes**		
	A	Enzymes and their classification ; Importance of Zn in nature, carbonic anhydrase, carboxypeptidase, alcohol dehydrogenase.		
	B	Biological nitrogen fixation (Nitrogenase) and abiological nitrogen fixation		
	C	tungsten containing formate dehydrogenase and tungsten bearing hyperthermophilic and thermophilic enzymes.		
	Unit 5	**Biologically Important Processes**		
	A	Photosynthetic electron transport chain, chlorophyll, PS-I and PS-II,		
	B	Vitamin B 12 coenzyme, its function and application in organic synthesis.		
	C	Availability of iron and iron toxicity.		
	Mode of examination	Theory		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1. S. J. Lippard & J. M. Berg. Principles of Bioorganic Chemistry; Panima Publ. Corpn. (2005). 2. E. -I. Ochiai. Bioinorganic Chemistry; An Introduction; Allyn and Bacon Inc. (1977).		
	Other References	1.M. N. Hughes. The Inorganic Chemistry of Biological Processes; Wiley (1981). 2.R. P. Hanzlik. Inorganic Aspects of Biological and Organic Chemistry; Academic Press (1976). 3.H. Kraatz & N. Metzler-Nolte (Eds.). Concepts and Models in Bioinorganic Chemistry; Wiley (2006). 4.Bertini; H. B. Gray; S. J. Dippard & J. S. Valentine; Bioinorganic Chemistry; Viva Books Pvt. Ltd. (2004). 5.A. W. Addison; W.R. Cullen; D. Dolphin & B.R. James (eds.). Biological Aspects of Inorganic Chemistry; John Wiley (1977).		

2.1 Template A1: Physical Chemistry-V (MCH239)

School: SBSR		**Batch : 2020-2022**
Program: M. Sc		**Current Academic Year: 2021-22**
Branch: Chemistry		**Semester:IV**
1	Course Code	MCH 239
2	Course Title	**PHYSICAL CHEMISTRY-V**
3	Credits	4.0.0
4	Contact Hours (L-T-P)	(4 0 0)
	Course Status	Compulsory
5	Course Objective	1. To provide the understanding of Quantum mechanical aspect of Band gap and Band theory in semi conductors. 2. To understand the various techniques for the preparation of nanomaterial and properties of nanomaterials. 3. To extend the concept of X-Ray diffraction, their generation and different experiments to study X-Ray diffraction. 4. To provide the understanding of physical aspects in Biological phenomenon. 5. To provide the indepth concept of polymers and their properties.
6	Course Outcomes	CO1: Direct and indirect band gap in semiconductors, types and analysis of p-n junctions. CO2: Students will be able to prepare nanomaterials and will be able to characterize their optical, electronic and structural properties. CO3:Students will be able to understand the generation of X-rays and diffraction patterns and will be able to refine the X-ray patterns. CO4: Students will be able to understand the energy Transformation and Thermodynamic principles and their applications in biological system. CO5: Student will be able to calculate the molecular weights of polymers using different techniques and will able to identify different physical and rheological properties of polymers. CO6: In depth knowledge of semiconductors, nanomaterials and polymers with application of X-rays, their generation and refinement of structure, application of physical phenomenons in biological system.
7	Course Description	Course emphasizing on the application part of Solid state chemistry and analysis of structure using X-Ray diffraction, materials chemistry, Biophysical aspects and applications and properties of polymers.
8	Outline syllabus	
	Unit 1	**Solid State Chemistry**
	A	Free electron theory of metals, Quantum mechanical treatment

		explaining the origin of band gaps, density of states, Band theory, Bloch theorem, Brillouin zones, effective mass of charge carriers,
	B	Semiconductors: Direct and indirect band gap semiconductors, hole concept, temperature dependence of mobility and electrical conductivity, free carrier concentration in intrinsic and extrinsic semiconductors, mass active law,
	C	Generation of carriers and their recombination in semiconductors. Types of junctions (metal-semiconductor, semiconductor-semiconductor, junctions in organic materials), Analysis of p-n junction including I-V characteristics.
	Unit 2	**Materials Chemistry**
	A	Definition of nanomaterials, various techniques for the preparation of nanomaterials, Thermodynamics and Kinetics of Nucleation, Thin Films and
	B	Langmuir-Blodgett films - Preparation techniques, evaporation/sputtering, chemical processes, MOCVD, sol-gel. Langmuir-Blodgett (LB) film growth techniques,
	C	photolithography, properties and applications of thin and LB films. Electronic structure and properties of nanomaterials, optical, electrical and magnetic properties, Chemical behaviour, applications of nanomaterials.
	Unit 3	**X-Ray Diffraction and Crystal Structure**
	A	Generation of X-rays**,** diffraction of X-rays by crystals, systematically absent reflections, multiplicities,
	B	X-ray diffraction experiments: the powder method- Bragg condition, Laue method, Bragg method and single crystal method, scattering of X-rays by atoms and a crystal,
	C	Patterson Synthesis, the Rietveld Refinement of $BaTiO_3$, ZnO and $BaSnO_3$, R-factor.
	Unit 4	**Biophysical Chemistry**
	A	Energy Transformation and Distribution of Energy, Thermodynamic principles in biological systems; Osmotic pressure, membrane equilibrium,
	B	muscular contraction and energy generation in mechanochemical system.Cell Membrane and Transport of Ions: Structure and functions of cell membrane.
	C	Active transport across cell membrane, irreversible thermodynamics treatment of membrane transport.
	Unit 5	**Polymers**
	A	Introduction, Classification of polymers, Concept of Mass and Number average molecular weights, Methods of Determining molecular weights (osmometry,
	B	diffusion and light scattering), Physical properties of polymers (glass

		transition temperature, crystalline melting point),		
	C	Rheological Properties, Biodegradable and Biomedical polymers, Liquid crystal polymers.		
	Mode of examination	Theory		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1.Polymer Chemistry, Billmayer 2. Polymer Chemistry, Gowarikar 3. Biological Thermodynamics,Donald T. Haynie, Cambridge. 4.Biophysical Chemistry, Vol. 1-3, C. R. Cantor & Schimmel 5. Biophysical Chemistry: Principles and Techniques by Upadhyay, Himalaya Publishing House 6.Introduction to Biophysical chemistry, R. Bruce Martin, McGraw-Hill, N 964. 7. Solid State Chemistry and its Applications(1984), A.R. West, John Wi nd Sons, Singapore 8. Introduction to Solids(1977), L.V. Azaroff, Tata McGraw-Hill, New Delhi 9. Solid State Chemistry(1992), L. Smart and E Moore, Chapman & Hall, Madras 10. Principles of Solid State(1993), H. V. Keer, Wiley Eastern 11. Instrumental methods of chemical analysis: Braun		

2.1 Template A1: Organic Chemistry-V (MCH240)

School: SBSR		**Batch : 2020-22**
Program: M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester: IV**
1	**Course No.**	**MCH240**
2	**Course Title**	**Organic Chemistry V**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Compulsory
5	Course Objective	1.To impart knowledge on synthesis of five and six- member heterocyclic compounds with two or more hetero atoms. 2.To familiarize with the synthesis of larger ring heterocyclic compounds. 3.To impart knowledge on synthesis of natural products. 4.To familiarize with structure determination and stereochemistry of terpenoids and carotenoids. 6.To understand the structure and significance of alkaloids.
6.	Course Outcomes	CO1: Understand the structure, properties, synthesis and reactions of five and six- member heterocyclic compounds with two or more hetero atoms. CO2: Propose syntheses and applications of heterocycles from the major classes. CO3:Describe the structure and synthesis of terpenoids and carotenoids. CO4:Formulate the synthesis of few important alkaloids. CO5:Identify medicinal properties of alkaloids. CO6: Acquire basic knowledge of natural product chemistry and understand the importance heterocycles in biological systems and in pharmaceuticals.
7	Course Description	This course will provide a concise introduction to heterocyclic chemistry. Emphasis will be given on the most important heterocyclic systems particularly five, and six-membered heterocyclic systems with 2- or 3-heteroatoms as well as fused heterocyclic systems. Chemical synthesis, properties, characteristics and applications of these systems will be discussed in detail. The course provides a basic knowledge of natural products chemistry with emphasis on terpenoids, carotenoids and alkaloids.
8	Outline syllabus	
	Unit 1	**Heterocycles I**
	A	Introduction**,** synthetic approaches, reactions and important applications of five membered heterocyclic compounds with two or three hetero atoms - imidazole, oxazoles,
	B	synthetic approaches, reactions and important applications of - thiazoles, oxadiazoles,
	C	synthetic approaches, reactions and important applications of -

		thiadiazoles, triazole.		
	Unit 2	**Heterocycles II**		
	A	Introduction**,** synthetic approaches, reactions and important applications of condensed five and six membered heterocycles with one hetero atom – indole,		
	B	synthetic approaches, reactions and important applications of – benzofuran, benzothiophene,		
	C	Synthetic approaches, reactions and important applications of – quinoline and isoquinoline.		
	Unit 3	**Heterocycles III**		
	A	Introduction**,** synthetic approaches, reactions and important applications of six membered heterocyclic compounds with two hetero atoms – pyridazine.		
	B	synthetic approaches, reactions and important applications of pyrimidine		
	C	synthetic approaches, reactions and important applications of pyrazine.		
	Unit 4	**Terpenoids and carotenoids**		
	A	Classification, nomenclature, occurrence, isolation, general methods of structure determination, isoprene rule. Structure determination and synthesis of the following representative molecules: Monoterpenoids - Citral, geraniol (acyclic), α-terpeneol, menthol (monocyclic). Sesquiterpenoids - Farnesol (acyclic), zingiberene (monocyclic), santonin (bicyclic), Diterpenoids - Phytol and abietic acid, β- carotene, lycopene and vitamin A.		
	B	Structure determination and synthesis of the following representative molecules: Sesquiterpenoids - Farnesol (acyclic), zingiberene (monocyclic), santonin (bicyclic),		
	C	Structure determination and synthesis of the following representative molecules: Diterpenoids - Phytol and abietic acid, β- carotene, lycopene and vitamin A.		
	Unit 5	**Alkaloids**		
	A	Definition, nomenclature and physiological action, occurrence, isolation, general methods of structure elucidation, degradation, classification based on nitrogen heterocyclic ring, role of alkaloids in plants.		
	B	Occurence, synthesis and structure elucidation of alkaloids – Reserpine		
	C	Occurence, synthesis and structure elucidation of alkaloids –morphine.		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
9	Text Book/s*	1.Heterocyclic Chemistry, T. L. Gilchrist. 2.An Introduction to the Chemistry of Heterocyclic compounds, R. M. Acheson. 3. Heterocylic chemistry, J. A. Joule & K. Mills. 4. Principles of Modern Heterocyclic Chemistry, A. Paquette. 5. Heterocyclic Chemistry, J. A. Joule & Smith. 6. Handbook of Heterocyclic Chemistry, A. R. Katritzky.		

		7.Natural Products : Chemistry and Biological significance, J. Mann, R. S. Davidson, J. B. Hobbs, D. V., Banthropde & J. B. Harborne. 8.Organic Chemistry, Vol-2, I. L. Finar
10	References	1.Stereoselective Synthesis: A Practical Approach, M. Nogrudi. 2.Rodd's Chemistry of Carbon Compounds, Ed. S. Coffey. 3. Chemistry, Biological and Pharmacological properties of Medicinal plants from the Americans, Ed. Kurt. Hostettmann, M. P. Gupta and A. Marston

2.1 Template A1: Inorganic Chemistry-VI (MCH241)

School: SBSR		**Batch : 2020-22**
Program:M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester:IV**
1	**Course Code**	**MCH241**
2	**Course Title**	**Inorganic Chemistry VI**
3	Credits	4
4	Contact hours	4-0-0
	Course Status	Compulsory
5	Course Objectives	1.Understand the importance of superconductors in engineering applications. 2.Relate the supramolecular role in ion detections. 3.Understand the chemistry of glasses and ceramics and their application in daily routine. 4.Understand the role of superconductors in catalysis. 5.Describe the technique used in applications of nanomaterials. 6.Understand the importance of nanomaterial based device in daily. routine.
6	Course Outcome	CO1:Understand the concept of molecular recognition in the application of supramolecules. CO2:Relate the applications of glass and ceramics on the basis of their structure. CO3:Explain the concept of superconductivity. CO4:Synthesis of nanomaterials. CO5:Identify the properties of nanomaterials and their applications in electronic applications. CO6:Gain knowledge about various advanced inorganic materials.
7	Course Description	The course is framed to give broad view of supramolecular, smart inorganic materials, superconductors and nanomaterials. Physicochemical properties and applications of nanomaterials have been covered in this paper.
8	Outline syllabus	
	Unit 1	**Supramolecular Chemistry**
	A	Concepts of Molecular recognition: Molecular receptors for different types of molecules including anionic substrates, design and synthesis of co-receptor molecules and multiple recognition
	B	Catenanes, Rotaxanes, Dendrimers and Supramolecular gels, Supramolecular reactivity in catalysis
	C	Transport processes and carrier design. Supramolecular devices. Some example of self-assembly in supramolecular chemistry
	Unit 2	**Inorganic Smart Materials**
	A	Structure of Glass and Ceramics: Ceramics crystal structures,density computations, silicate ceramics
	B	Glass ceramics.Refractories with reference to preparation, Properties and

		applications.		
	C	fibre reinforced Composites, microscopic composites, preparation procedure, special properties and applications		
	Unit 3	**Superconductors**		
	A	Inorganic semiconductors, Electrical, magnetic, thermal and optical properties of superconductors,.		
	B	Metallic bonds High temperature superconductors Structural features of cuprate superconductors:1-2-3 and 2-1-4 cuprates.		
	C	Electrical and magnetic properties of superconductors		
	Unit 4	**Nanomaterials**		
	A	Definition of nanomaterials, fullerenes**,** carbon nanotubes, graphene. Discovery of C_{60}, Superconductivity in C_{60}, Alkali doped C_{60}.		
	B	Carbon nanotubes - Synthesis of Single walled carbon nanotubes. Synthesis methods - Arc discharge, Laser Abalation, Low temperature method, Chemical vapour deposition. Growth mechanisms on CNT.		
	C	Structure and characterization techniques. Surface area measurement, determination of size and textural studies of nanotubes.		
	Unit 5	**Physiochemical Properties and Applications of Nanomaterials**		
	A	Reactivity, effect of size and shape on nanocrystal reactivity, agglomeration and sintering, dispersibility and chemical stability in solution, surface modification of metallic and semiconductor nanoparticles, nanofabrication and nanomanipulation.		
	B	Magnetism in nanomaterials, Doping, functionalizing nanotube.		
	C	Applications of Graphene, CNTs and Fullerenes – sensing, organic transistor, odour sensor, electronics and optoelectronics and photovoltaics.		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
9	Textbook/s*	1.Timp.G. , Ed.Nanotechnology, Springer-Verlag, N. Y 2.Supramolecular Chemistry by Jonathan W Steed; Jerry L. Atwood.		
10	Other References	1.Keer, H.V. Principles of the Solid State, Wiley Eastern Ltd., New Delhi. 2.West, A. R., Solid State Chemistry and its applications, John Wiley and Sons. 3.Supramolecular Chemistry: Concept and Perspective by Jean Marie Lehn 4.Mitchell, B. S. – An introduction to material engineering and Science.Wiley interscience.		

2.1 Template A1: PHYSICAL CHEMISTRY VI (MCH 242)

School: SBSR		**Batch : 2020-2022**
Program: M. Sc		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester: 04**
1	**Course Code**	**MCH 242**
2	**Course Title**	**PHYSICAL CHEMISTRY VI**
3	Credits	4
4	Contact Hours (L-T-P)	(3 1 0)
	Course Status	Compulsory
5	Course Objective	6. To provide the understanding of photophysical and photochemical processes of atoms and diatomic molecules. 7. To understand various nonradiative relaxation processes. 8. To get familiar with high energy radiation with matter, radiation dosimetry and flash photolysis. 9. To understand the meaning, scope, laws of irreversible thermodynamics. 10. To provide information about various laws, parameters, and equations related to transport phenomenon. 11. To provide the conceptual knowledge of molecular and advanced photochemistry; radiation chemistry, dosimetry, and photolysis; irreversible thermodynamics and transport phenomenon.
6	Course Outcomes	CO1: To understand various photophysical and photochemical processes of atoms and diatomic molecules upon irradiation. CO2: To study the various radiationless relaxation pathways. CO3: To learn about mechanism of interaction of high energy radiation with matter; radiation dosimetry and principle and application of flash photolysis. CO4: To understand the fundamental meaning, scope, and laws of irreversible thermodynamics. CO5: To get familiarize with different parameters and laws related to transport phenomenon. CO6: To study molecular and advanced photochemistry; radiation chemistry, dosimetry, and photolysis; irreversible thermodynamics and transport phenomenon.
7	Course Description	Course emphasize on the basic concepts of molecular and advanced photochemistry; radiation chemistry, dosimetry, and photolysis; irreversible thermodynamics and transport phenomenon.
8	Outline syllabus	
	Unit 1	**Molecular photochemistry**
	A	Introduction-primary photophysical process of atoms and diatomic molecules, the absorption and emission of light - spectroscopic notations, state mixing, spin-orbit coupling and spin forbidden radiative transitions,
	B	Absorption complexes, Franck-Condon principle, selection rules, laws of

		photochemical equivalence. Radiative transitions-classical model of radiative transitions. Transitions between states (chemical, classical and quantum dynamics, vibronic states).
	C	Potential energy surfaces; transitions between potential energy surfaces. Jablonski diagram, Fluorescence, phosphorescence, photosensitization, photosynthesis, and chemiluminescence.
	Unit 2	**Advanced photochemistry**
	A	Wave mechanical interpretation of radiationless transitions between states, factors influencing the rate of vibrational relaxation. Fluorescence quenching: collisional quenching, Stern-Volmer equation, concentration quenching, quenching by excimer and exciplex emissio
	B	Energy transfer: Theory of radiationless energy transfer and energy transfer by electron exchange. Fluorescence resonance energy transfer between photoexcited donor and acceptor systems and dexter energy transfer.
	C	The Perrin formulation. Triplet-triplet, triplet-singlet, singlet triplet energy transfer. Multiphoton energy transfer processes, reversible energy transfer.
	Unit 3	**Radiation Chemistry, Dosimetry and Photolysis: An overview**
	A	G-value. The mechanism of interaction of high energy radiation with matter, Photoelectric effect, Compton effect, Pair production, total absorption co-efficient, excitation and ionization, Stopping power and linear energy transfer.
	B	Radiation dosimetry: Radiation dose and its measurement, standard free air chamber method, chemical dosimeter (Frick's Dosimeter). Short lived intermediates (ions, excited molecules, free radicals: Various mechanisms of their formation and energy transfer processes)..
	C	Flash photolysis: Principle and its applications. Radiolysis of water and aqueous solutions. Radiolysis of molecules of biological interest (carbohydrates, amino acids, peptides, and nucleic acids).
	Unit 4	**Irreversible thermodynamics**
	A	Meaning and scope of irreversible thermodynamics, Thermodynamic criteria for non-equilibrium states, Phenomenological laws- Linear laws, Gibbs equation,
	B	Onsager's reciprocal relations, Entropy production- specific examples of entropy production, Non-equilibrium stationary states,
	C	Prigogine's principle of maximum entropy production, Coupled phenomena. Some important applications.
	Unit 5	**Transport phenomena**
	A	Diffusion coefficients, Fick's first and second laws, relation between flux and viscosity,
	B	relation between diffusion coefficient and mean free path, relation

		between thermal conductivity/viscosity and mean free path of a perfect gas, Einstein relation,		
	C	Nernst-Einstein equation, Stokes-Einstein equation, Einstein-Smoluchowski equation.		
	Mode of examination	Theory		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1. Turro, N. J. Modern Molecular Photochemistry Univ. Science Books (1991). 2. Gilbert, A. & Baggot, J. Essentials of Molecular Photochemistry Blackwell Scientific (1990). 3. Sood, D.D., Reddy, A.V.R. and Ramamoorthy, N., "Fundamentals of Radiochemistry", IANCAS, BARC, Mumbai. 4. Mukherjee, K.K., "Fundamentals of Photochemistry", New Age International Pvt. Ltd., New Delhi. 5. Lakowicz, J.R., "Principles of Fluorescence Spectroscopy", Plenum Press, New York. 6. Wishart, J.F. and Nocera, D.G., "Photochemistry and Radiation Chemistry", Oxford University Press, USA. 7. Friedlander, G., Kennedy J.W., Miller, E.S. and Macais, J.M., "Nuclear and Radiochemistry", John Wiley and Sons, Inc. New York. 8. Atkin's Physical Chemistry, P. Atkins & Julio de Paula, Oxford University Press 9. Introduction to Thermodynamics of Irreversible Processes by I. Prigogine, Interscience 10. Fundamentals of Photochemistry, K. K. Rohatgi-Mukherjee. 11. Katchalsky, A. & Curren, P. F. *Non Equilibrium Thermodynamics in Biophysics* Harvard University Press: Cambridge (1965). 12. Kalidas, C. & Sangaranarayanan, M.V. Non-Equilibrium Thermodynamics: Principles & Applications, Macmillan India Ltd. (2002).		

2.1 Template A1: Organic Chemistry-VI (MCH243)

School: SBSR		**Batch 2020-22**
Program: M.Sc.		**Current Academic Year : 2020-22**
Branch : Chemistry		**Semester IV**
1	**Course No.**	**MCH243**
2	**Course Title**	**Organic Chemistry VI**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course status	Compulsory
5	Course Objective	1.To provide a comprehensive introduction to biochemistry. 2.To learn the chemistry of enzymes, structures of nucleic acids, proteins and carbohydrates. 3.To know the chemistry of selected steroids, cholesterol and hormones. 4.To familiarize the chemistry and structure of oxytocin. 5.To know the kinetics of enzymes. 6.To understand the chemistry of antibiotics.
6	Course Outcomes	CO1:To introduce structure and functions of carbohydrates and their derivatives. CO2:Understand the structure, function, and folding of proteins. CO3:Analyze the double helical structure of DNA and its replication, RNA and transcription. CO4:Learn kinetics of enzyme catalyzed reactions and enzyme inhibition. CO5:Convert cholesterol to progesterone, estrone and testosterone and structure elucidation of cholesterol. CO6:Acquire knowledge of molecular structure and interactions present in proteins, nucleic acids and carbohydrates and enzymes, the organization and working principles of various components present in living cell.
7	Course Description	The course is designed to give provide an ability to assess the significance of fundamental chemical properties on biomolecular structure, understanding of the connection between biomolecular structure and function, acquire knowledge of chemical synthesis of biomolecules and the chemical reactions of biomolecules.
8	Outline Syllabus	
	Unit 1	**Carbohydrates**
	A	Conformation of monosaccharides, structure and functions of important derivatives of monosaccharides like glycosides, deoxy sugars, myoinositol, amino sugars.
	B	N-acetylneuraminic acid, sialic acid disaccharides and polysaccharides. Structural polysaccharides - cellulose and chitin. Storage polysaccharides- starch and glycogen. Structure and biological functions of glucosaminoglycans or mucopolysaccharides.

	C	Carbohydrates of glycoprotiens and glycolipids. Role of sugars in biological recognition. Blood group substances. Ascorbic acid.		
	Unit 2	**Amino acids and Proteins**		
	A	Chemical and enzymatic hydrolysis of proteins to peptides, amino acid sequencing. Secondary structure of protein, forces responsible for holding of secondary structures. A- helix, β-sheets, super secondary structure, triple helix structure of collagen. Tertiary structure of protein-folding and domain structure. Quaternary structure		
	B	Amino acid metabolism- degradation and biosynthesis of amino acids, sequence determination: chemical/ enzymatic/ mass spectral, racemization/ detection		
	C	Chemistry of oxytocin and tryptophan releasing hormone (TRH).		
	Unit 3	**Nucleic Acids**		
	A	Introduction, chemical and enzymatic hydrolysis of nucleic acids, Structure physical and chemical properties of the heterocyclic bases – Adenine, Guanine. Cytosine, Uracil and Thiamine.		
	B	Structure and synthesis of mono and poly – nucleosides and nucleotides. Deoxyribose nucleic acid (DNA): Primary, secondary, tertiary structure of DNA.Structure of RNA. Types of RNA – mRNA, rRNA and tRNA.		
	C	The chemical basis for heredity, an overview of replication of DNA, transcription, translation and genetic code.		
	Unit 4	**Enzymes**		
	A	Introduction and historical perspective, chemical and biological catalysis, remarkable properties of enzymes like catalytic power, specificity and regulation.		
	B	Nomenclature and classification, extraction and purification. Fischer's lock and key and Koshland's induced fit hypothesis, concept and identification of active site by the use of inhibitors, affinity labeling and enzyme modification by site-directed mutagenesis.		
	C	Enzyme kinetics, Michaelis–Menten and Lineweaver Burk plots, reversible and irreversible inhibition, mechanism of enzyme action		
	Unit 5	**Steroids and Hormones**		
	A	Occurrence, nomenclature, Diel's hydrocarbon and stereochemistry.		
	B	Isolation, structure determination and synthesis of Cholesterol, bile acids		
	C	Androsterone, testosterone, estrone, progesterone, vitamin D		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text Book/s*	1.A.L. Lehninger, Principles of Biochemistry, CBS Publishers, Delhi. 2.I.L. Finar Volume II.		
	Other references	1.D. Voet, J.G. Voet & CW Pratt, Fundamentals of Biochemistry, John Wiley & Sons, New York. 2.H.R. Mahler and E.H. Cordes, Biological Chemistry, 2nd Edition, Harper and Row Pub., New York. 3.T.C. Bruice and S. Bentkovic, Bioorganic Mechanisms, Vol. I & II, W. A. Benjamin, New York.		

2.1 Template A1: Medicinal Chemistry (MCE203)

School: SBSR		**Batch : 2020-22**
Program:M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester:IV**
1	**Course No.**	**MCE203**
2	**Course Title**	**Medicinal Chemistry**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Elective
5	Course Objective	1.To provide a comprehensive introduction to Pharmaceutical Chemistry. 2.To introduce the Quantitative structure activity relationship. 3.To introduce the software used in drug designing. 4.To explain the process of pharmacology. 5.To introduce the chemistry of antineoplastic drugs. 6.To throw light on the chemistry of Anti-HIV Drugs and AIDS and antibiotics.
6.	Course Outcomes	CO1:Explain concept of Quantitative Structure Activity Relationship. CO2:Understand the process of pharmacokinetic and pharmacodynamics. CO3:Elucidate the mode of action of Antineoplastic drugs. CO4:Explain the chemistry and mode of action of Anti-HIV and AIDS drugs. CO5:Explain the chemistry and mode of action of NSAID drugs and review the chemistry of Antibiotic drugs. CO6:Have a thorough grounding in Pharmaceutical Chemistry and basic knowledge in drug designing.
7	Course Description	The course is emphasises on physical interactions and chemical reactions and their mechanisms as applied to biological systems, how drugs are discovered and developed, classified, how they get to their site of action, what happens when they reach the site of action in their interaction with receptors, enzymes, and DNA. The approaches discussed are those used in the pharmaceutical industry and elsewhere for the discovery of new drugs.
8	Outline Syllabus	
	Unit 1	**Drug Design and Development**
	A	Procedures followed in drug design, concept of lead compound and lead modification
	B	structure-activity relationship (SAR), Quantitative structure activity relationship (QSAR). History and development of QSAR. Physicochemical parameters: lipophilicity, Hydrophobicity, Electronic effect Steric factors, Hansch equation (Mathematical derivations of equations excluded).
	C	Computer aided drug design. Software used in drug design.
	Unit 2	**Pharmacology**

	A	Pharmacokinetics: various modes of administration of drug, distribution, metabolism (biotransformation) and drug excretion		
	B	pharmacodynamic: Concepts of drug receptors interactions		
	C	Definition of the following medicinal terms: Pharmacon, pharmacophore, soft drug, prodrug, half-life, efficiency, LD50, ED50, therapeutic index, drug toxicity, drug addiction, spurious drugs, misbranded drugs, adulterated drugs, pharmacopoeia		
	Unit 3	**Antineoplastic Agents**		
	A	Introduction, cancer chemotherapy, special problems		
	B	Role of alkylating agents and antimetabolites in treatment of cancer. Mode of action of mechlorethamine, cyclophosphamide, 5-Fluorouracil.		
	C	Recent development in cancer chemotherapy.		
	Unit 4	**Anti-HIV Drugs and NSAIDs**		
	A	Basic facts about HIV & AIDS, Structure of HIV cell, Anti HIV drugs and their classification		
	B	NSAIDS & Mechanism of Action:Asprin		
	C	NSAID-Induced Side Effects		
	Unit 5	**Antibiotics**		
	A	Introduction, classification of antibiotics, β-lactam antibiotics & their mode of action - Amoxicillin, Chloramphenicol, Cephalosporin		
	B	Tetracycline antibiotics & their mode of action, Aminoglycoside antibiotics & their mode of action - Streptomycin.		
	C	Macrolide antibiotics & their Mode of action - erythromycin		
	Mode of examination	Theory/Jury/Practical/Viva		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1.Strategies for Organic Drug Synthesis and Design, D. Lednicer, John Wiley & Sons Ltd. 2A Text Book of Medicinal Chemistry, Vol-I and Vol-II, Surendra N. Pandeya, SG Publishers. 3.An Introduction to Drug Design, S.S. Pandeya and J. R. Dimmock, New Age International Publishers. 4.Medicinal Chemistry, Ashutosh Kar, New Age International Publishers. 5.Goodman and Gilman's Pharmacological Basis of Therapeutics, McGraw-Hill.		
	Other References	1.Introduction to Medicinal Chemistry, A. Gringauge, Wiley-VCH. 2. Wilson and Gisvold's Text book of Organic Medicinal and Pharmaceutical Chemistry, Edited by J.N. Delgado and W. A. Remers, J.B. Lipincott Company. 3.The Organic Chemistry of Drug Design and Drug Action, R.B. Silverman, Academic Press. 4.Burger's Medicinal Chemistry and Drug Discovery, Vol. I-V, Edited by M.E. Wolff, John Wiley & Sons Ltd		

2.1 Template A1: Chemistry of Nanomaterials (MCE204)

School:SBSR		**Batch:2020-22**
Program:M.Sc.		**Current Academic Year: 2020-22**
Branch:Chemistry		**Semester:IV**
1	**Course Code**	**MCE204**
2	**Course Title**	**Chemistry of Nanomaterials**
3	Credits	4
4	Contact Hours (L-T-P)	4-0-0
	Course Status	Elective
5	Course Objective	1.Teach the advanced methods towards the synthesis of functional materials. 2.Teach the advanced methods towards the synthesis of high-quality thin films. 3.Teach the mechanical and magnetic behaviour of functional materials. 4.Teach the basics and phenomenon associated with the electrical and optical behavior. 5.Teach modern spectroscopic and microscopic methods towards the characterization of functional materials. 6. To understand the novel materials from synthetic, analysis and application perspectives.
6	Course Outcomes	CO1:Formulate the synthetic methods towards preparation of novel materials. CO2:Prepare the mechanistic pathway towards facile synthesis of thin films. CO3:Understand the diverse magnetic behaviour of materials CO4:Understand the various electro-optical phenomenon of the materials. CO5:Characterize the materials via spectroscopic and microscopic tools. CO6:Understand the advanced synthetic perspectives along with physical properties and the concept of Auger and X-ray Photoelectron Spectroscopy.
7	Course Description	The elective course on Chemistry of Materials aims to teach the modern and advanced methods of synthesis, characterization and properties of novel materials.
8	Outline syllabus	
	Unit 1	**Synthesis Methods: Physicochemical Techniques**
	A	Preparation of materials by Ball milling, Attrition and Vibration milling, Cluster compounds, Preparation of nano particles, Preparation of nanostructured polymers/Conducting polymers, composites.
	B	Chemical precipitation and co-precipitation, Wet chemical methods, Metal crystals by reduction, Sol-gel synthesis
	C	Microemulsions or reverse micelles, Hydrothermal & Solvothermal synthesis, Thermolysis routes, Microwave heating synthesis,

		Electrochemical synthesis.		
	Unit 2	**Synthesis Methods: Deposition Techniques**		
	A	Physical Vapor Deposition; mass evaporation rate; evaporators, e-beam, reactive evaporation, ion beam assisted deposition, Sputtering techniques		
	B	Chemical Vapor Deposition - reaction chemistry and thermodynamics of CVD		
	C	Thermal CVD, laser & plasma enhanced CVD, Pyrolytic synthesis.		
	Unit 3	**Unit 3: Properties: Mechanical and Magnetic**		
	A	**S**tress Strain diagram for different engineering materials, Ductile and brittle material, Tensile strength, Hardness, Impact strength		
	B	Fracture (Types and Ductile to brittle transition), Fatigue, Creep, Factors affecting mechanical properties		
	C	Classification of magnetic materials, Diamagnetism, Paramagnetism, Langevin theory of dia- and paramagnetism, Ferromagnetism, Antiferromagnetism, Ferrimagnetism, Structure of Ferrite.		
	Unit 4	**Properties: Electrical and Optical**		
	A	Dielectric Materials: Basic concepts: complex permittivity, dielectric loss factor, polarization, mechanism of polarization, classification of dielectrics-frequency dependence of dielectric constant		
	B	Ferroelectricity, Piezoelectricity, pyro-electric states, transition temperature, polarization catastrophe, antiferroelectricity, ferro electric domains.		
	C	Optical Properties: Refractive index and dispersion, Transmission, Reflection and absorption of light, Optical material for UV and IR, Optical anisotropic, Non-linear optical crystals, Photoluminescence		
	Unit 5	**Structural Analysis**		
	A	UV-visible, FT-IR, Raman and Atomic absorption spectroscopy; X-ray diffraction		
	B	Glancing angle and wide angle, Debye-Scherer formula, Dislocation density, Micro strain		
	C	AUGER Spectroscopy and X-ray photoelectron spectroscopy (XPS)		
	Mode of examination	Theory		
	Weightage Distribution	CA	MTE	ETE
		30%	20%	50%
	Text book/s*	1.Characterization of materials (Vol. 1 and 2) by E.N. Kaufmann, John Wiley and Sons. 2.Structure and Properties of Materials', Volume III,by R. M., Rose Shepard L. A., Wulff J.,4th Edition, John Wiley, 1984		
	Other References	1.Pradeep T., "NANO the Essential, understanding Nanoscience and Nanotechnology". TataMcGraw-Hill Publishing Company Limited, 2007. 2.Charles P. Poole Jr. "Introduction to Nanotechnology", John Willey & Sons, 2003		

2.3 A3: Syllabus of CCU401

School: SBSR		**Batch :2020- 2022**		
Program: M.Sc.		**Current Academic Year: 2020-22**		
Branch: Chemistry		**Semester: II**		
1	Course Code	CCU401		
2	Course Title	Community Connect		
3	Credits	2		
4	Contact Hours (L-T-P)	2-0-0		
	Course Status	Compulsory		
5	Course Objective	1. To expose our students to different social issues faced by the people in different sections of society. 2. To connect their class-room learning with problem solving skills in real life scenario.		
6	Course Outcomes	After completion of this course students will be able to: 1. Recognise social problems prevailing in different sections of society and finding the solution in sustainable manner. 2. Get practical exposure of all round development which complements their class room learning 3. These activities will add value to students, faculty members, school and university.		
7	Course Description	In this mode, students will make survey, analyze data and will extract results out of it to correlate with their theoretical knowledge. E.g. Crops and animals, land holding, labour problems, medical problems of animals and humans, savage and sanitation situation, waste management etc.		
8	Outline syllabus			
	Unit 1	**Introduction to the Topic**		
	Unit 2	**Drafting the questionairre**		
	Unit 3	**Survey**		
	Unit 4	**Data collection, Discussions and result interpretation**		
	Unit 5	**Report writing and Presentation**		
	Mode of examination	Presentation and Viva		
	Weightage Distribution	CA	MTE	ETE
		60%	0%	40%
	Text book/s*	-		
	Other References	The entries in the list should be in alphabetical order.		

		Journal article Hamburger, C.: Quasimonotonicity, regularity and duality for nonlinear systems of partial differential equations. Ann. Mat. Pura Appl. 169, 321–354 (1995) Article by DOI Sajti, C.L., Georgio, S., Khodorkovsky, V., Marine, W.: New nanohybrid materials for biophotonics. Appl. Phys. A (2007). doi:10.1007/s00339-007-4137-z Book Geddes, K.O., Czapor, S.R., Labahn, G.: Algorithms for Computer Algebra. Kluwer, Boston (1992) Book chapter Broy, M.: Software engineering — from auxiliary to key technologies. In: Broy, M., Denert, E. (eds.) Software Pioneers, pp. 10–13. Springer, Heidelberg (2002) Online document Cartwright, J.: Big stars have weather too. IOP Publishing PhysicsWeb. http://physicsweb.org/articles/news/11/6/16/1 (2007). Accessed 26 June 2007 Always use the standard abbreviation of a journal's name according to the ISSN List of Title Word Abbreviations, see www.issn.org/2-22661-LTWA-online.php For authors using End Note, Springer provides an output style that supports the formatting of in-text citations and reference list. End Note style (zip, 2 kB)

2.3 A3: Syllabus of Dissertation A

School: SBSR		**Batch :2020- 2022**		
Program: M.Sc.		**Current Academic Year: 2020-22**		
Branch: Chemistry		**Semester: V**		
1	Course Code	MCH276		
2	Course Title	Dissertation A		
3	Credits	2		
4	Contact Hours (L-T-P)	0-0-6		
	Course Status	Compulsory/Elective		
5	Course Objective	1.To enhance the practical knowledge and result analysis skills. 2.To enable the students experience a real-life problem solving under the supervision of faculty members. 3.To prepare the students perform functions that demand higher competence in national/international organizations. 4.To train the students in scientific research. 5.To help the students find meaning in life by broadening their field of vision. 6.Develop deep knowledge of a specific area of specialization by literature search.		
6	Course Outcomes	CO1: Able to do literature search, develop deeper interest / inquisitiveness in chemistry and interdisciplinary subjects. CO2: Able to prepare stock solutions, buffers etc . CO3: Understand the basics of chemistry and become familiar with qualitative and qualitative estimations. CO4: Able to understand the chemistry of reactions. CO5: Able to analyse the results and understand the chemical reactions involved. CO6: Enhance the practical skills.		
7	Course Description	This course provides the applied knowledge of chemistry and gives confidence and a solid foundation for future learning.		
8	Outline syllabus			
	Unit 1	**Introduction of subject / Literature search**		
	Unit 2	**Concept building and Study designing**		
	Unit 3	**Experimentation / Standardization of techniques**		
	Unit 4	**Data collection, Discussions and result interpretation**		
	Unit 5	**Report writing**		
	Mode of examination	Presentation and Viva		
	Weightage Distribution	CA	MTE	ETE
		60%	0%	40%
	Text book/s*	-		
	Other References	Pubmed Search (NCBI) Review and research articles of Indexed Journals		

2.3 A3: Syllabus of Dissertation B

School: SBSR		**Batch :2020- 2022**		
Program: M.Sc.		**Current Academic Year: 2020-22**		
Branch: Chemistry		**Semester: VI**		
1	Course Code	MCH275		
2	Course Title	Dissertation B		
3	Credits	6		
4	Contact Hours (L-T-P)	0-0-12		
	Course Status	Compulsory/Elective		
5	Course Objective	1.To enhance the practical knowledge and result analysis skills. 2.To enable the students experience a real-life problem solving under the supervision of faculty members. 3.To prepare the students perform functions that demand higher competence in national/international organizations. 4.To train the students in scientific research. 5.Develop research/ experimentation skills as well as enhancing project writing and oral presentation skills 6.Inculcate team spirit and time management.		
6	Course Outcomes	CO1: Able to use lab instruments independently. CO2:Cultivate the understanding of problem, study design, methodology/ experimentation, significance of reproducibility of results. CO3:Understanding of ethics of science and research for supporting higher studies. CO4:Learn effective project organizational skills along with discussions, result interpretation and paper writing. CO5: Able to analyse the results and understand the chemical reactions involved. CO6: Enhance the practical skills.		
7	Course Description	This course will help to develop knowledge and research skills applicable to a career in chemistry.		
8	Outline syllabus			
	Unit 1	**Introduction of subject/ literature search**		
	Unit 2	**Concept building and study design**		
	Unit 3	**Experimentation/ Standardization of techniques**		
	Unit 4	**Data collection, Discussions and result interpretation**		
	Unit 5	**Report writing**		
	Weightage Distribution	CA	MTE	ETE
		60%	0%	40%
	Text book/s*	-		
	Other References	Pubmed Search (NCBI) Review and research articals of Indexed Journals		